



Title	Tris(2-methoxyphenyl)bismuthane as a Dehydrating Agent with High Template Ability : an Efficient Single-step Synthesis of Macrocyclic Diesters from Diacid Anhydrides and Glycols
Author(s)	Ogawa, Takuji; Yoshikawa, Akemi; Wada, Hidemi et al.
Citation	JOURNAL OF THE CHEMICAL SOCIETY-CHEMICAL COMMUNICATIONS. 1995, 14, p. 1407-1408
Version Type	VoR
URL	https://hdl.handle.net/11094/3008
rights	
Note	

The University of Osaka Institutional Knowledge Archive : OUKA

<https://ir.library.osaka-u.ac.jp/>

The University of Osaka

Tris(2-methoxyphenyl)bismuthane as a Dehydrating Agent with High Template Ability: an Efficient Single-step Synthesis of Macrocyclic Diesters from Diacid Anhydrides and Glycols

Takuji Ogawa,*^a Akemi Yoshikawa,^a Hidemi Wada,^a Chie Ogawa,^a Noboru Ono^a and Hitomi Suzuki*^b

^a Department of Chemistry, Faculty of Science, Ehime University, Bunkyo-cho, Matsuyama 790, Japan

^b Department of Chemistry, Faculty of Science, Kyoto University, Kitashirakawa, Sakyo-ku, Kyoto 606-01, Japan

Tris(2-methoxyphenyl)bismuthane works as both a mild dehydrating agent and a good template for macrocyclic ester synthesis; prolonged heating of the bismuthane with a dicarboxylic acid anhydride and a glycol in toluene under reflux afforded the corresponding macrocyclic 1:1 diester in moderate to good yields, together with a small amount of a macrocyclic 2:2 tetraester.

In spite of being a member of pnictogen family, trivalent organobismuth exhibits a weak Lewis acid nature due to the valence extension ability characteristic of heavy non-metallic elements. There are many reports which describe the coordination of a bismuth atom with the oxygen or nitrogen atoms of ether and amino compounds.¹ However, such a prominent and specific interacting ability of the bismuth atom with lighter heteroatoms has apparently not been utilized in organic synthesis to date. We report herein that tris(2-methoxyphenyl)bismuthane **1** works as both a mild dehydrating agent² and as a good template for the macrocyclic ester synthesis due to its excellent coordinating ability; when heated with compound **1** in toluene under reflux, dicarboxylic acid anhydride and glycol smoothly coupled together to give the corresponding macrocyclic diester in a higher yield than previously reported.³

A typical preparation of macrocyclic diesters is as follows: A mixture of phthalic anhydride **2a** (740 mg, 5.0 mmol), tetraethylene glycol **3** ($n = 3$) (970 mg, 5.0 mmol), bismuthane **1** (1.06 g, 2.0 mmol) and dry toluene (10 cm³) was heated under reflux for 2 d. A white deposit was filtered off and the filtrate was evaporated under reduced pressure to leave a pale yellow oil, which was purified by silica gel column chromatography with 5% MeOH-CH₂Cl₂ as eluent. A 1:1 diester **6a** ($n = 2$) was obtained as the major product from the early eluates and a 2:2 tetraester **7a** ($n = 2$) as the minor product from the later eluates. The isolated yields of compounds **6a** ($n = 2$) and **7a**




($n = 2$) were 61 and 3% respectively. The results are summarized in Table 1.

Although MS inspection of the crude products revealed the formation of cyclic 2:2 tetraesters **7** in every case, only **7a** ($n = 2, 3$), **7c** ($n = 2$) and **7d** ($n = 2$) could be successfully isolated in a pure form. All products were identified by spectral and analytical data.[†] The macrocyclic 1:1 diester and 2:2 tetraester structures were confirmed unambiguously by X-ray analyses of **6a** ($n = 2$) and **7a** ($n = 3$), see Figs. 1 and 2.[‡]

These types of macrocyclic compounds were previously prepared in low yields (1–30%) from the reaction between diacid dichlorides and glycols by the high dilution method (*ca.* 60 mmol dm⁻³, 3 d).³ By using bismuthane **1** as a template, the corresponding macrocyclic diesters could be obtained in much better yields (13–83%) at higher concentrations (*ca.* 500 mmol dm⁻³).

The initial stage of the macrocyclic formation would be an ordinary reaction between acid anhydride **2** and glycol **3** to give a monoester **4**, which then reacts with the bismuthane **1** to form a key intermediate **5**. In this intermediate, the oxygen atoms of the ethylene glycol moiety coordinate to the bismuth atom,^{1a} and subsequently the bismuthane-mediated dehydration occurs intramolecularly to produce the macrocyclic diester **6** (Scheme 1). High-yield formation of macrocyclic compounds even at high concentrations are possible due to the dual role of the bismuthane **1** as the dehydrating agent and the template under the reaction conditions employed.

Table 1 Preparation of macrocyclic di- and tetra-esters

Acid anhydride	X	Polyethylene glycol 3 n	Yield (%) ^a (mp/°C)	
			6	7
2a		2	18 (123–125)	13 (127–128)
		3	61 (94–96)	3 (111–112)
		4	13 (oil)	— ^b
2b		3	60 (oil)	— ^b
		4	56 (oil)	— ^b
2c	—(CH ₂) ₃ —	2	20 (oil)	9 (oil)
		3	47 (oil)	— ^b
		4	83 (oil)	— ^b
2d	—(CH ₂) ₂ —	2	17 (59–62)	8 (oil)
		3	48 (oil)	— ^b
		4	65 (oil)	— ^b
2e		3	trace (oil) ^c	— ^b

^a Yields refer to the isolated compounds and were not optimized. ^b Not isolated. ^c Polymeric products were formed.

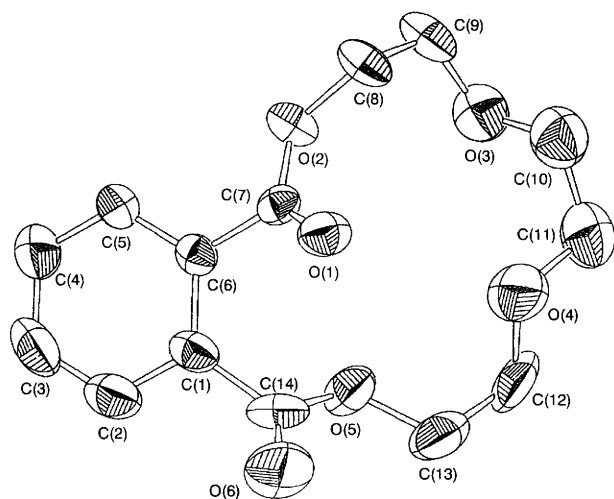


Fig. 1 An ORTEP drawing of macrocyclic 1:1 diester **6a** ($n = 2$)

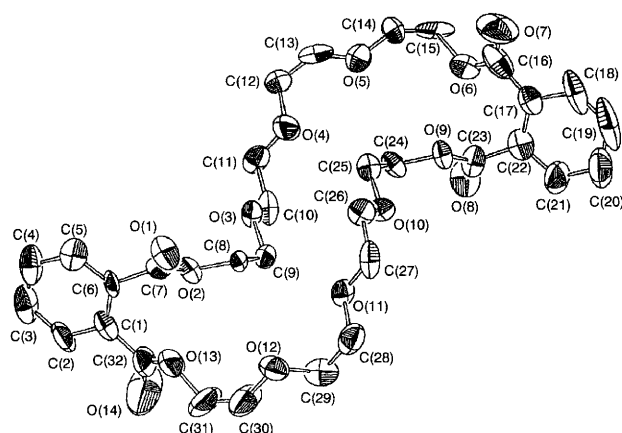
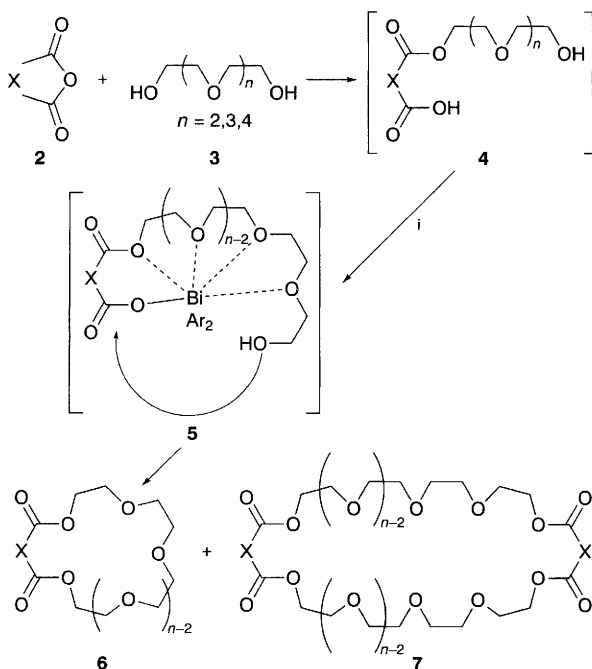


Fig. 2 An ORTEP drawing of macrocyclic 2:2 tetraester **7a** ($n = 3$)



Scheme 1 Reagents and conditions: i, bismuthane **1**, PhMe, reflux, 2 d

We thank the Ministry of Education, Science and Culture of Japan for financial support of this work by a Grant-in-Aid for Science and Research (No. 05804038).

Received, 10th March 1995; Com. 5/01510C

Footnotes

† Selected spectroscopic and analytical data for compound **6a** ($n = 2$): mp 123–125 °C; δ_{H} (J/Hz) 3.66 (4 H, s), 3.84 (4 H, t, J 4.8), 4.50 (4 H, t, J 4.8) and 7.28–7.75 (4 H, m); m/z (CI) 281 (8%, $M + 1$), 237 (8), 193 (54), and 149 (100) (Found: C, 59.7; H, 5.8. $\text{C}_{14}\text{H}_{16}\text{O}_6$ requires C, 60.0; H, 5.8%). For **7a** ($n = 2$): mp 127–128 °C; δ_{H} (J/Hz) 3.68 (8 H, s), 3.78 (8 H, t, J 5.0), 4.44 (8 H, t, J 5.0) and 7.51–7.74 (8 H, m); m/z (CI) 561 (100%, $M + 1$). **6a** ($n = 3$): mp 94–96 °C; δ_{H} (J/Hz) 3.66 (8 H, bs), 3.86 (4 H, t, J 6), 4.48 (4 H, t, J 6), 7.53 (2 H, dd, J 3.3 and 5.8) and 7.73 (2 H, dd, J 3.3 and 5.8); δ_{C} 61.69, 64.72, 68.92, 70.59, 128.98, 131.09 and 167.38; ν_{max} (KBr) cm^{-1} 1710, 1590, 1570 and 1060; m/z (CI) 325 (100%, $M + 1$) (Found: C, 58.6; H, 6.4. $\text{C}_{16}\text{H}_{20}\text{O}_7$ requires C, 59.3; H, 6.2%). For **6a** ($n = 4$): oil; δ_{H} (J/Hz) 3.64 (4 H, bs), 3.68 (8 H, bs), 3.84 (4 H, t, J 4.9), 4.49 (4 H, t, J 4.9), 7.53 (2 H, dd, J 3.3 and 5.8) and 7.74 (2 H, dd, J 3.3 and 5.8); δ_{C} 65.07, 68.96, 70.67, 70.79, 128.99, 131.02 and 131.96; m/z (CI) 369 (4%, $M + 1$), 237 (40) and 194 (100). (Found: C, 58.7; H, 6.7. $\text{C}_{18}\text{H}_{24}\text{O}_8$ requires C, 58.7; H, 6.6%). For **6b** ($n = 3$): oil; δ_{H} (J/Hz) 1.4–2.0 (8 H, m), 2.77 (2 H, m), 3.64 (12 H, bs) and 4.23 (4 H, t, J 6); m/z (CI) 331 (18%, $M + 1$) and 155 (100). For **6b** ($n = 4$): δ_{H} (J/Hz) 1.4–2.0 (8 H, m), 2.85 (2 H, bs), 3.66 (16 H, bs) and 4.21 (4 H, t, J 4); m/z (CI) 375 (10%, $M + 1$). **6c** ($n = 2$): δ_{H} (J/Hz) 2.07 (2 H, quintet, J 7.2), 2.42 (4 H, quintet, J 7.2), 3.66 (8 H, bs) and 4.27 (4 H, t, J 6); m/z (CI) 247 (100%, $M + 1$). For **6c** ($n = 3$): oil [lit.³ bp 155–157 °C/0.65 Torr], δ_{H} (J/Hz) 1.94 (2 H, m), 2.40 (4 H, t, J 5), 3.64 (12 H, bs) and 4.24 (4 H, t, J 4.8); ν_{max} (NaCl) cm^{-1} 1710 and 1040; m/z (CI) 291 (100%, $M + 1$). For **6c** ($n = 4$): δ_{H} (J/Hz) 2.0 (2 H, quintet, J 7), 2.4 (4 H, quintet, J 7), 3.66 (16 H, bs) and 4.25 (4 H, t, J 4.8); m/z (CI) 335 (100%, $M + 1$). For **6d** ($n = 2$): δ_{H} (J/Hz) 2.66 (4 H, s), 3.66 (8 H, bs) and 4.27 (4 H, t, J 4.8); m/z (CI) 233 (100%, $M + 1$). For **6d** ($n = 3$): δ_{H} (J/Hz) 2.66 (4 H, s), 3.65 (8 H, bs), 3.68 (4 H, t, J 5) and 4.24 (4 H, t, J 5); ν_{max} (NaCl) cm^{-1} 1710 and 1040; m/z (CI) 277 (100%, $M + 1$). For **6d** ($n = 4$): δ_{H} (J/Hz) 2.66 (4 H, s), 3.66 (16 H, bs) and 4.27 (4 H, t, J 4); m/z (CI) 321 (100%, $M + 1$).

‡ Crystal data for **6a** ($n = 2$). $\text{C}_{14}\text{H}_{16}\text{O}_6$, $M = 280.28$. Orthorhombic prisms, $a = 14.429(4)$, $b = 11.662(9)$, $c = 8.294(3)$ Å, $V = 1396(2)$ Å³, space group $Pna2_1$, $Z = 4$, $D_c = 1.334$ g cm⁻³. Intensity data were recorded on a Rigaku AFC5R diffractometer with graphite-monochromated Mo-K α radiation and a 12 kW rotating anode generator. Crystal dimensions 0.350 \times 0.180 \times 0.200 mm. Scans of $(0.68 + 0.30 \tan \theta)^\circ$ were made at a speed of 16.0° min⁻¹ (in omega). The final cycle of full-matrix least-squares refinement was based on 751 observed reflections [$I > 3.00 \sigma(I)$] and 244 variable parameters and converged with unweighted and weighted agreement factors of $R = 0.042$ and $R_w = 0.043$. The maximum and minimum peaks on the final difference Fourier map corresponded to 0.20 and -0.17 e Å⁻³, respectively. For **7a** ($n = 3$). $\text{C}_{32}\text{H}_{40}\text{O}_{14}$, $M = 648.66$. Monoclinic prisms, $a = 14.404(5)$, $b = 15.436(4)$, $c = 15.982(4)$ Å, $\beta = 114.66(2)^\circ$, $V = 3230(2)$ Å³, space group $P2_1/n$, $Z = 4$, $D_c = 1.334$ g cm⁻³. Scans of $(1.26 + 0.30 \tan \theta)^\circ$ were made at a speed of 16.0° min⁻¹ (in omega). The final cycle of full-matrix least-squares refinement was based on 1702 observed reflections [$I > 3.00 \sigma(I)$] and 415 variable parameters and converged with unweighted and weighted agreement factors of $R = 0.076$ and $R_w = 0.088$. Crystal dimension 0.450 \times 0.450 \times 0.080 mm. The maximum and minimum peaks on the final difference Fourier map corresponded to 0.36 and -0.27 e Å⁻³, respectively. Atomic coordinates, bond lengths and angles and thermal parameters have been deposited at the Cambridge Crystallographic Data Centre. See Information for Authors, Issue No. 1.

References

- (a) Coordination of the polyethylene glycol moieties to a bismuth atom has been established by X-ray crystallographic studies; R. D. Rogers, A. H. Bond and S. Aguinaga, *J. Am. Chem. Soc.*, 1992, **114**, 2960. As for recent examples of the coordination of lighter heteroatom elements to bismuth, see (b) M. Wieber, H. G. Rudling and C. Burschka, *Z. Anorg. Allg. Chem.*, 1980, **470**, 171; M. Dräger and B. M. Schmidt, *J. Organomet. Chem.*, 1985, **290**, 133; C. A. Stewart, J. C. Calabrese and A. J. Arduengo, III, *J. Am. Chem. Soc.*, 1985, **107**, 3397; M. Veith and B. Bertsch, *Z. Anorg. Allg. Chem.*, 1988, **557**, 7; F. Huber, M. Domagala and H. Preut, *Acta Crystallogr., Sect. C*, 1988, **44**, 828; M. A. Matchett, M. Y. Chiang and W. E. Buhro, *Inorg. Chem.*, 1990, **29**, 358; H. Suzuki, T. Murafuji and N. Azuma, *J. Chem. Soc., Perkin Trans. 1*, 1992, 1593; K. Whitmire, J. C. Hutchinson, A. L. McKnight and C. M. Jones, *J. Chem. Soc., Chem. Commun.*, 1992, 1021; Y. Yamamoto, X. Chen and K. Akiba, *J. Am. Chem. Soc.*, 1992, **114**, 7906.
- T. Ogawa, T. Hikasa, T. Ikegami, N. Ono and H. Suzuki, *J. Chem. Soc., Perkin Trans. 1*, 1994, 3473.
- M. D. Thompson, J. S. Bradshaw, S. F. Nielsen, C. T. Bishop, F. T. Cox, P. E. Fore, G. E. Maas, R. M. Izatt and J. J. Christensen, *Tetrahedron*, 1977, **33**, 3317.