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Structure of Slag (IV)†

—Behaviour of Amphoteric Metal Ions in Slag—

Nobuya IWAMOTO *

Abstract

Behaviour of amphoteric ions such as aluminium and titanium in slag have been considered with the results from various state analysis means. It has been emphasized that the behaviour of titanium ion in slag still remained unclear when compared with that of aluminium ion.

1. Introduction

As previously reported,1)–3) the concept of the basicity of slag still remains unclear despite the efforts of many investigators.

As logic, the problem can be resolved if we could know the activity of $O^{2-}$ ion in slag. Recently Wagner says that the evaluation of emf measurement of cells involving slags is not recommended even if only an approximate measurement of the basicity of a slag is wanted.4)

In 1953, Tomlinson said that the solubility of $CO_2$ gas in slag is probably proportional to the activity of $O^{2-}$ ion in slag.5)

Pearce measured the solubility of $CO_2$ gas in glasses of $Na_2O-SiO_2$ and $Na_2O-B_2O_3$ systems.6),7) Likewise, Holmquist studied the change of the solubility of $SO_3$ gas in glass of $Na_2O-SiO_2$ system.8) These sorts of studies seem worthy to obtain an effective information about slag structure.

In the welding division, it is especially important to know distribution coefficients of $Mn/(MnO)$ and $Si/(SiO_2)$ between metal and slag. Generally, it is now discussed from the viewpoint that there is interrelation between the distribution coefficients and the basicity of slag.9)–11)

As previously emphasized,1),2) the role of the additives such as $Al_2O_3$, $TiO_2$ and $CaF_2$ on the basicity of slag is unknown, although these components are frequently used for flux in welding.

Accordingly, it is the purpose of this report to search new index explaining the interrelation.

In the glass technology, intermediate oxides including the ions of Be, Zn, Pb, Fe, Co, Ni, Ti, and Al have been paid our attention on the coordinated state in glasses.12)

Until now, we were continuously concerned with these problems. Therefore, I would like to give a certain direction to elucidate the behaviours of aluminium and titanium ions in slag.

2. Problem on aluminium ion in glass and slag

Schairer and Bowen originally measured the refractive index of glasses of $Na_2O-SiO_2-Al_2O_3$ system.13) From the isofracts obtained, it was determined that the bent occurs around the line connecting $SiO_2$ and $Na_2O/AI_2O_3 = 1$ as shown in Fig. 1. Galant found same tendency, that is, the break of refractive index occurred when the amounts of $Al_2O_3$ and $Na_2O$ contained in the glass became equivalent.14) He simply said that aluminium can be four- or six- coordination according to the ratio of $Al_2O_3$ and $R_2O$ in alkali aluminosilicate glasses.

Isard measured the electrical conductivities of soda aluminosilicate glasses.15) He found the occurrence of minima of activation energy for electrical conduction

† Received on Jan. 8, 1976
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when the Al/Na ratio equaled unity. The result is shown in Fig. 2.
Kozakevitch measured viscosity of CaO-Al₂O₃-SiO₂ system ranging widely of 1800 - 2100°C. As an example, iso-viscosity diagram at 1900°C is shown in Fig. 3. He found the similar tendency which produced the viscosity crest near the line CaO/Al₂O₃ = 1. The change was attributed to the difference between four- and six-coordination of aluminium ion in slag.

Kolesova studied the infrared absorption spectra of alkali aluminosilicate glasses and concluded that the absorption spectra resulted in the region near 760 cm⁻¹ of alkali aluminosilicate glasses commenced to develop two maxima with the increase of Al₂O₃ content of glass.²³

Saksena performed the infrared absorption studies of some silicate minerals and he paid his attention on the problem of Al-O bonds. As his conclusion, it was presented that it was impossible to distinguish four- or six-coordination of aluminium ions from the infrared absorption study.

Tarte classified the main absorption of XO₄ and XO₆ as follows.

### Table 1 Characteristic frequency of some coordinated groups

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<th>Cation X</th>
<th>Absorption range (cm⁻¹)</th>
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<tr>
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<td>XO₄ tetrahedra</td>
</tr>
<tr>
<td></td>
<td>XO₆ octahedra</td>
</tr>
<tr>
<td>Ti⁴⁺</td>
<td>Condensed</td>
</tr>
<tr>
<td>Al</td>
<td>900 - 700</td>
</tr>
<tr>
<td></td>
<td>800 - 690</td>
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<td></td>
<td>800 - 650</td>
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In the complex aluminosilicate, the force constants of the Al-O bonds of AlO₄ tetrahedra are rather lower than those of Si-O bonds, but generally the vibrations give a composite absorption band including both Al-O-Si and Si-O-Si stretching vibrations in the 900 - 1100 cm⁻¹ region, which shifts to lower frequencies with increasing aluminium for silicon substitution.

Day and Rindone said that the four-fold coordination state of the aluminium ions which do not precipitate in the tetrahedra network could occasionally be surrounded by five, six, or seven oxygen ions and have an average coordination number of six. Accordingly, they said that the fact, which the infrared absorption spectra of glasses having an Al/Na ratio greater than unity did not contain an absorption band, could be assigned specifically to AlO₄ groups.

Tarte summarized that the picture is incomplete as six-fold coordination by infrared absorption spectroscopy for several reasons:
1) The characteristic frequency range is generally not so well delimited as in the case of four-fold coordi-
nation.

2) These vibrations are generally located in the low frequency region, the experimental study of which is difficult.

Besides the difficulties inherent to the study of the groups with high coordination numbers, the application of infrared spectroscopy to structural studies of slag suffers from more general drawbacks such as:

1) Lack of sensitivity, especially due to the broadness of the bands.

2) Difficulties in the correct assignment of the bands: in some cases, it is quite impossible to choose between several interpretations. For these reasons, it became advisable to obtain low frequency Raman bands using more powerful and monochromatic excitation laser sources.

Our results of IR and Raman spectra are compared in Figs. 4 and 5. From these spectra, it will be understood that the views above described are correct.

The Lorenz-Lorentz equation is used to calculate the molar refraction of each specimen from the density and refractive index. As the additive rule is assumed to hold, the partial molar refraction of the components added can be deduced. Molecular refraction for visible light is a measure of the deformation or distortion of the electron shells in the structure units.

Accordingly, it becomes possible to estimate the change of metallic cations coordinating in slag from the information. This subject was treated by Faján in detail. The application to glasses were carried out by three groups until now.

Safford and Silverman, and Day and Rindone treated the coordination problem of aluminium ion in glasses. However they failed to certify the existence of aluminium ions as six-fold state because of the experimental inaccuracies of the values of density and refractive index.

Turnbull and Lawrence studied the effect of TiO₂ addition to the systems, Na₂O-2SiO₂-xTiO₂ and Na₂O-4SiO₂-xTiO₂ on the partial molar refactivity of TiO₂. As an example, the effect of TiO₂ content on total refraction of a Na₂O-2SiO₂-xTiO₂ glass is shown in Fig. 6. From the result, it will be seen that there never occurs the abrupt change when the ratio TiO₂/Na₂O becomes more than unity. It was concluded that
TiO₂ by itself is not a glassformer but in the presence of silicon may enter into the network structure.

It has been well known that the information on the emission X-ray peak shift is effective to determine a change in the vaquence of an element.

Day summarized peak displacements of 0.04 to 0.06°(2θ) and 0.08 to 0.12°(2θ) for compounds having aluminium ions in fourfold and six-fold coordination respectively. His result is shown in Fig. 7. Brindlay and McKinstry noticed that X-ray peak displacement is suffered a considerable influence due to the kind of anion surrounding an aluminium ion. Experimental results obtained by us are given in Table 2. From this, it will be understood that aluminium ions in the Na₂O-SiO₂-Al₂O₃ and CaO-SiO₂-Al₂O₃ systems take four- or six-fold coordination depending on MO/Al₂O₃ ratio of slag composition from the comparison with the iso-total oxygen density curves are shown on the slags of CaO-SiO₂-Al₂O₃ system. It will be anticipated that similar tendency of the bend taking place around the line connecting SiO₂ apex and CaO/Al₂O₃ = 1.

3. Problem on titanium ion in glass and slag

In the welding, TiO₂ is widely used as additive for the welding flux because it gives flux to have proper viscosity characteristic.

On the other hand, it has been said that in the blast furnace operation with the slag containing titanium ion troublesome problem occurred because of highly viscous slag formation. These contradictory results are still unresolved.

Hellbrügge and Endel showed the contrary additive effect of TiO₂ and SiO₂ on the viscosity of Na₂O-2SiO₂ glass. The former decreases the viscosity coefficient, but the latter increases. From these results, it is
difficult to recognize the similar action in glass or slag between TiO₂ and SiO₂.

Frohberg and Schenck have investigated in detail on the viscosity coefficient of CaO-SiO₂-TiO₂ and Al₂O₃-SiO₂-TiO₂ system. The followings are their conclusions. It was noticed that SiO₂ content has great influence on the viscosity in Al₂O₃-SiO₂-TiO₂ system, but viscosity does not change noticeably in the case of low TiO₂ content. On the other hand, the isoviscosity lines are parallel to the variation of Al₂O₃ in the case of CaO-Al₂O₃-TiO₂ system. CaO content gives additional effect to the influence of TiO₂. They have summarized that TiO₂ generally decreases viscosity of the slag which contained CaO, SiO₂, and Al₂O₃.

Secondary, the results of electrical conductivity measurements will be briefly described. With increasing TiO₂ content in Na₂O-SiO₂-TiO₂ system, electrical conductivity decreases, but the value increases in CaO-SiO₂-TiO₂ system. It seems as if Ti⁴⁺ ion makes Ca²⁺ ion free from the silicate network. However, in FeO-SiO₂ system, electrical conductivity decreases with increasing of SiO₂, but increases with the addition of TiO₂.³³,³⁴

Turnbull and Lawrence concluded that the role of TiO₂ in silica glass resembled to that of crystals in which the coordination number is six from the measurement of density and refractive index.¹²

On the contrary, Rao presented the results that the Ti⁴⁺ ion is a network-former in its own right and is capable of taking part in the network in tetrahedral and octahedral coordination, depending on conditions from the various measurements such as oxygen packing, expansivity, viscosity and tensile strength.³⁵

Whereas, Rao said that the coordination of titanium would be changed from six to five in the transformation of BaTiO₃ and PbTiO₃.

From the viewpoint of glass-ceramics, many investigators took note of TiO₂ addition to various glasses. Mackenzie and Brown compared octahedral site preference energy for titanium ion in aluminosilicate with the published values by McClure and Dunitz and Orgel. From this they concluded that octahedral site preference energy of titanium ion is lower than Al³⁺, therefore, it is in predominantly tetrahedral site. Kondratev mentioned that titanium begins to enter the aluminosilicate structure in the form of [TiO₄] units, when the amount of TiO₂ introduced into the glass exceeds the amount of aluminum oxide in octahedral coordination.³⁹ Further, during repeated heating titanium tends to supplement its oxygen deficiency by some displacement of the atoms around the [TiO₄] units, which are converted into deformed [TiO₆] units. According to Galant, oxygen ions from alkali or alkaline earth oxides available in TiO₂ are sufficient for formation of [TiO₄] polyhedrons, and therefore, it can be assumed that titanium retains the coordination number of six in the glass.⁴⁰

Our result shown in Fig. 5 explained that the band at 280 and 770 cm⁻¹ became more obvious which were probably attributed to the intrinsic character of rutile TiO₂ with increasing TiO₂ content.²²,²³

Probably TiO₄ and TiO₆ units will be formed in slag of K₂O-SiO₂-TiO₂ system. X-ray peak shift study, refractive index and X-ray absorption measurements are described in our papers.²¹,²²

4. Summary

The influence of the coordination state of cations in slag on the physical and chemical properties of slag is important matter.

Furthermore, detailed studies such as gas solubility, surface property and so on are wanted.

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