



Title	Optical properties of substituted phthalocyanine rare-earth metal complexes
Author(s)	Yoshino, K. ; Lee, S. B. ; Sonoda, T. et al.
Citation	Journal of Applied Physics. 2000, 88(12), p. 7137-7143
Version Type	VoR
URL	https://hdl.handle.net/11094/75852
rights	
Note	

The University of Osaka Institutional Knowledge Archive : OUKA

<https://ir.library.osaka-u.ac.jp/>

The University of Osaka

Optical properties of substituted phthalocyanine rare-earth metal complexes

Cite as: Journal of Applied Physics **88**, 7137 (2000); <https://doi.org/10.1063/1.1316050>

Submitted: 17 February 2000 . Accepted: 14 August 2000 . Published Online: 28 November 2000

K. Yoshino, S. B. Lee, T. Sonoda, H. Kawagishi, R. Hidayat, K. Nakayama, M. Ozaki, K. Ban, K. Nishizawa, K. Ohta, and H. Shirai



View Online



Export Citation

ARTICLES YOU MAY BE INTERESTED IN

[Electronic structure of phthalocyanines: Theoretical investigation of the optical properties of phthalocyanine monomers, dimers, and crystals](#)

The Journal of Chemical Physics **92**, 1228 (1990); <https://doi.org/10.1063/1.458131>

[Phosphorescence and Fluorescence of Phthalocyanines](#)

The Journal of Chemical Physics **55**, 4131 (1971); <https://doi.org/10.1063/1.1676714>

[Charge transport in lithium phthalocyanine](#)

The Journal of Chemical Physics **104**, 5048 (1996); <https://doi.org/10.1063/1.471134>





Lock-in Amplifiers

Zurich
Instruments

Watch the Video 

Optical properties of substituted phthalocyanine rare-earth metal complexes

K. Yoshino^{a)} and S. B. Lee^{b)}

Department of Electronic Engineering, Faculty of Engineering, Osaka University, 2-1 Yamada-Oka, Suita, Osaka 565-0871 and Institute for Fundamental Chemistry, 34-4 Takano-nishihiraki-cho, Sakyo-ku, Kyoto 606-8103, Japan

T. Sonoda, H. Kawagishi, R. Hidayat, K. Nakayama, and M. Ozaki

Department of Electronic Engineering, Faculty of Engineering, Osaka University, 2-1 Yamada-Oka, Suita, Osaka 565-0871, Japan

K. Ban, K. Nishizawa, K. Ohta, and H. Shirai

Department of Functional Polymer Science, Shinshu University, Ueda 386-8567, Japan

(Received 17 February 2000; accepted for publication 14 August 2000)

Comparative study of optical properties of alkylthio-group-substituted phthalocyanine rare-earth metal sandwich complexes $[(C_nS)_8Pc]_2M$, $M=Eu, Lu, Tb$ is presented. Photoluminescence and photoconductivity of $[(C_nS)_8Pc]_2M$ complex is very weak. Two photoluminescence bands were observed at around 400–650 and 720–800 nm in chloroform solution corresponding to the Soret and Q bands in the absorption spectra, respectively. However, the emission from Eu^{3+} ion (as well as Tb^{3+}) was not found compared with other Eu complexes because the $5d$ levels of the Eu^{3+} ion lie higher than the triplet level of the ligand. The significant enhancement of the photoconductivity of $[(C_{16}S)_8Pc]_2M$ after C_{60} doping is reported. The photoconductivity is positive at the low electric field in the ohmic regime while it becomes negative at the high electric field upon photoexcitation with strongly absorbed light. The negative photoconductivity is attributed to space-charge effects. The mechanism of photoluminescence and photoconductivity are discussed by taking the electronic energy schemes of phthalocyanine ligands and lanthanide ion and C_{60} into consideration. © 2000 American Institute of Physics. [S0021-8979(00)05522-5]

I. INTRODUCTION

Quasione-dimensional nature of the columnar structure of discotic liquid crystals along which the fast carrier transport takes place stimulated much interest from both scientific and practical viewpoints. Extremely large hole mobility was observed in the discotic phase.^{1–3} Discotic liquid crystals have a potential technological importance due to their semi-conducting properties combined with their processability. When the discotic liquid crystal is heated to mesophase it self-organizes in one-dimensional discotic columnar structure and this columnar structure is preserved upon cooling to a solid state.⁴ High intracolumnar mobility of $1.7 \times 10^{-5} \text{ m}^2/\text{Vs}$ in discotic hexagonal columnar mesophase D_h of bis[octakis(dodecyloxy)phthalocyaninato]lutetium(III) at 90 °C measured by the contactless time-resolved microwave conductivity technique was reported.³ It should be mentioned that by designing the molecular structure of the central core parts of discotic liquid crystal by changing the central rare-earth metal ion or by varying the peripheral side groups the various functionalities are also expected. For ex-

ample, the alkylthio-substituted phthalocyanines have higher conductivity in the mesophase than alkoxy-substituted phthalocyanines.

Most of the discotic liquid crystals absorb light only in the UV region. The sensibilization by doping with visible light absorbing dyes is needed for such application as xerography and imaging devices. The discotic liquid crystals based on substituted metal phthalocyanine complexes strongly absorb visible light and, therefore, seem to be very interesting.

However, the photoluminescent and photoconductive properties of substituted phthalocyanine rare-earth metal complexes, which are very important for understanding the intramolecular and intermolecular charge transfer processes in this material and also for application in optoelectronics, have not been studied yet in detail, especially in the visible range. Only recently two reports appeared about the study of the photovoltaic effect in rare-earth bisphthalocyanine complex⁵ and liquid crystalline phthalocyanine.⁶

In this article a comparative study of optical properties of substituted phthalocyanine rare-earth metal complexes $[(C_{16}S)_8Pc]_2M$ ($M=Eu, Tb, \text{ and } Lu$) is presented. The mechanisms of photoluminescence (PL) and photoconductivity (PC) are discussed by taking the electronic energy schemes of phthalocyanine ligands and lanthanide (III) ion into consideration. The enhancement of photoconductivity after C_{60} doping at low electric field and negative photoconductivity at high electric field will also be discussed.

^{a)}Author to whom correspondence should be addressed; electronic mail: yoshino@ele.eng.osaka-u.ac.jp

^{b)}On leave from Heat Physics Department of Uzbek Academy of Science, Tashkent, Uzbekistan; present address: Dept. of Physics, University of Utah, 115 S. 1400 E Rm. 201, SLC, Utah 84112.

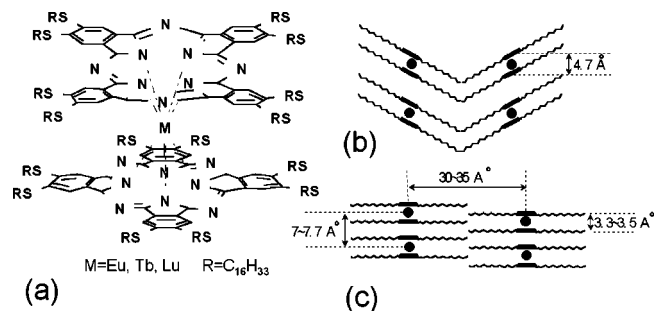


FIG. 1. (a) The molecular structure of bis[octakis(alkylthio)phthalocyaninato]lanthanide(III) complex $[(C_nS)_8Pc]_2M$. (b), (c) Schematic representations of the structural arrangement of $[(C_nS)_8Pc]_2M$ complexes in crystalline solid phase (b) and discotic liquid-crystalline phase (c).

II. EXPERIMENT

$[(C_{16}S)_8Pc]_2M$ ($M = Lu, Tb$, and Eu) complex was synthesized by a modified method of Belarbi's manner.⁷ Double decker molecular structure of the alkylthio-group-substituted phthalocyanine rare-earth metal sandwich complex, bis[octakis(hexadecylthio)-phthalocyaninato]lanthanide(III), is shown in Fig. 1(a). Two phthalocyanine macrocycles are 45° staggered with respect to each other. Schematic representations of the structural arrangement of phthalocyanine dimer in crystalline solid phase and discotic liquid-crystalline phase are shown in Figs. 1(b) and 1(c), respectively. The detailed synthetic procedures and mesomorphic properties will be reported separately.⁸

Absorption and photoluminescence spectra measurements were carried out utilizing a Hewlett-Packard 8452 A diode array spectrophotometer and a Hitachi F-2000 fluorescence spectrophotometer, respectively. For the study of fluorescence emission in broad spectral range a third harmonic generation light (350 nm) of Nd:YAG laser ($1.064 \mu m$) of 10 ns in pulse width was used as an excitation light source and a Hamamatsu photonic multichannel spectral analyzer (PMA-11). Time resolved PL was studied utilizing the femtosecond laser system (Tsunami, Spectra-Physics) with pulse width of 80 fs, wavelength of 390 nm, and detected using a Hamamatsu streak scope camera with temporal resolution of about 15 ps.

PC was measured by the conventional method as already reported by the authors.⁹ All measurements were carried out in vacuum of about 10^{-5} Torr after 1d pumping to avoid the oxygen effect.

III. RESULTS

A. Absorption and photoluminescence

A study of the solution system provides us with an insight into the intrinsic nature of the charge carrier photogeneration process. Figure 2(a) shows the optical absorption spectra of $[(C_{16}S)_8Pc]_2Eu$, $[(C_{16}S)_8Pc]_2Tb$, and $[(C_{16}S)_8Pc]_2Lu$ in chloroform solution. In this figure, two sets of absorption bands at 200–450 and 600–750 nm, which are related to the famous *B* (Soret) and *Q* bands of phthalocyanine, respectively, are clearly observed. The other three UV absorption peaks at 320, 244, and 210–215 nm could be

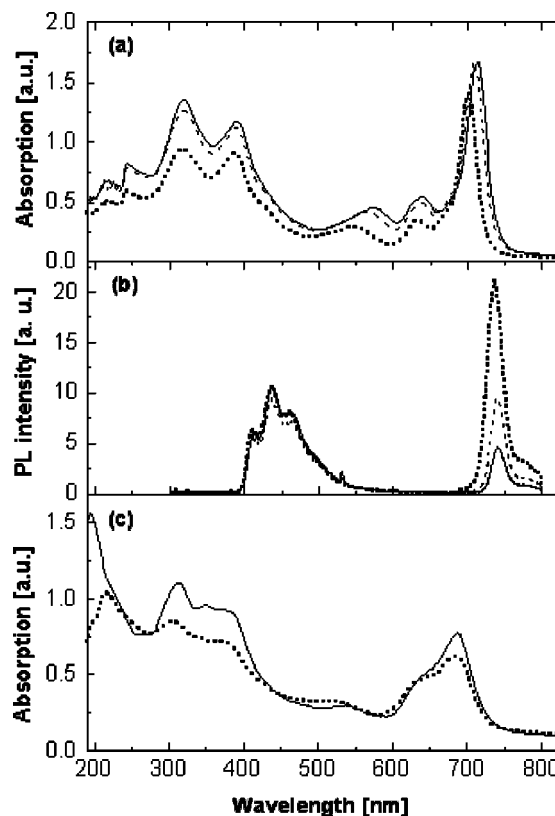


FIG. 2. (a) Optical absorption and (b) photoluminescence spectra of $[(C_{16}S)_8Pc]_2Eu$ (solid line), $[(C_{16}S)_8Pc]_2Tb$ (dashed line), and $[(C_{16}S)_8Pc]_2Lu$ (dotted line) in chloroform solution (5×10^{-7} M). PL excitation was third harmonic of Nd-YAG, 350 nm, 21 mJ/pulse, 10 ns ($r = 0.4$ cm, 4.2 kW/cm^2). (c) Optical absorption spectra of thin film of $[(C_{16}S)_8Pc]_2Eu$ (solid line) and $[(C_{16}S)_8Pc]_2Lu$ (dotted line).

ascribed to the so-called *N*, *L*, and *C* bands, respectively.¹⁰ The small bathochromic shift of *Q* absorption bands with the increase of atomic number of rare-earth metal ions can be detected while Soret and high energy absorption peaks were not changed. Schechtman and Spicer first noticed in their early study of phthalocyanines that the energies and strengths of UV transitions are relatively insensitive to substitutions of the central metal atom.¹¹ Several absorption peaks of Eu^{3+} and Tb^{3+} ions in visible range are too weak to be observed in the absorption spectrum.¹²

Figure 2(b) shows the spectral dependence of photoluminescence of $[(C_{16}S)_8Pc]_2M$ ($M = Eu, Tb$ and Lu). Two PL bands were detected simultaneously at around 400–650 and 720–800 nm in chloroform solution upon photoexcitation with a third harmonic generation light (350 nm) of Nd:YAG laser of 10 ns in pulse width.

Figure 2(c) shows the optical absorption spectra of thin film of $[(C_{16}S)_8Pc]_2Eu$ and $[(C_{16}S)_8Pc]_2Lu$. The main features of the individual molecules of the $[(C_{16}S)_8Pc]_2M$ complex are retained in thin film of this molecular material in the spectral range studied.

We have carried out more detailed measurement of emission and excitation spectra of substituted phthalocyanine rare-earth metal sandwich complexes. Figures 3(a) and 3(b) show the emission and excitation spectra of $[(C_{16}S)_8Pc]_2Eu$ in chloroform solution at room temperature. The high energy

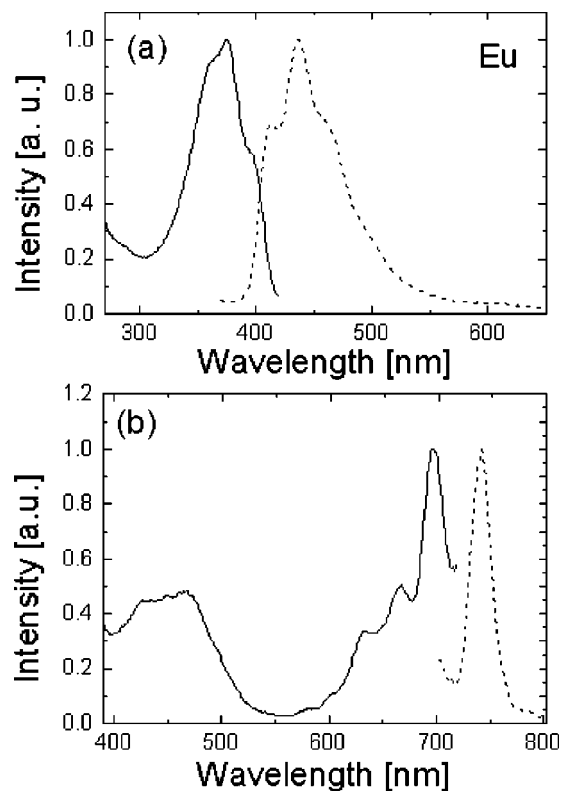


FIG. 3. PL emission (dotted line) and excitation (solid line) spectra of $[(C_{16}S)_8Pc]_2Eu$ in chloroform solution in visible range (a) and in near IR region (b).

PL band is vibrationally broadened with the first peak at about 3.0 eV and associated vibronic bands separated by 0.16–0.17 eV. The low energy PL band consists of a sharp peak at 740 nm and shoulder at about 765 nm. The position and relative intensities of these PL bands correlate well with the corresponding *B* and *Q* absorption bands, respectively. The redshift of low energy PL peaks was observed upon going from Lu to Eu complexes in correspondence with the shift of the *Q* band in absorption spectra.

The time-resolved PL spectroscopy at room temperature by photoexcitation in the high absorption band (Soret band) and observation of the decay of high energy PL band has shown that the decay curve can be fitted by a double exponent with $t_1 = 0.73$ ns and $t_2 = 2.2$ ns.

The strong PL quenching effect in chloroform solution of $[(C_{16}S)_8Pc]_2M$ with increasing concentration due to aggregation effect¹³ was observed together with changes of relative intensity of the two bands (Fig. 4).

The PL was completely suppressed in film. However, the PL still can be detected in solid disordered phase. Figure 5 shows the PL spectra of $[(C_{16}S)_8Pc]_2M$ infiltrated into nanoporous disordered opal matrices made of 180 nm silica spherical particles. The PL quenching effect was partially suppressed by destruction of the columnar structure. The enhancement of the PL intensity can even be observed by multiple light scattering if silica particles of proper sizes would be used.

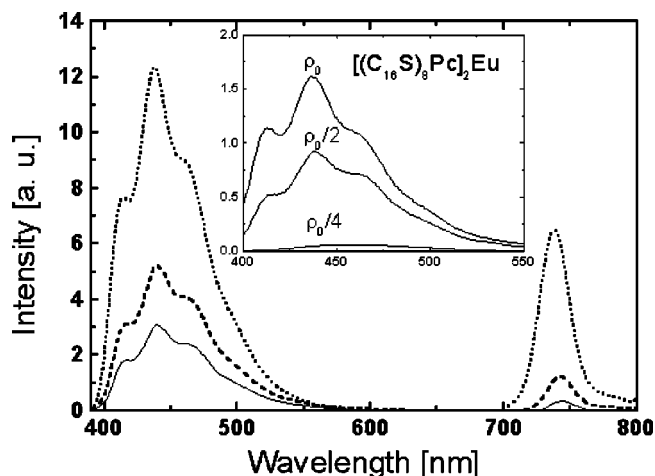


FIG. 4. Spectral dependence of PL of $[(C_{16}S)_8Pc]_2M$ ($M=Eu$ —solid line, Tb —dashed line, and Lu —dotted line) in chloroform solution (2×10^{-5} M). Photoexcitation was third harmonic light of Nd-YAG, 350 nm, 21 mJ/pulse, 10 ns. The inset shows the strong PL quenching effect for $[(C_{16}S)_8Pc]_2Eu$ for different concentration ($\rho_0 = 2 \times 10^{-5}$ M).

B. Photoconductivity

In Fig. 6 we show the spectral dependence of photoconductivity of $[(C_{16}S)_8Pc]_2Eu$ before and after C_{60} doping and also absorption spectrum of the undoped $[(C_{16}S)_8Pc]_2Eu$ thin film for comparison.

The sample was made by casting the chloroform/toluene solution of $[(C_{16}S)_8Pc]_2Eu$ and C_{60} on quartz substrates with comb-like patterned interdigital Au/Cr electrodes of 100 μm width separated by a distance of 100 μm . Such in-plane electrode geometry allows us to diminish the possible effect of photoinjection from electrodes and to study the intrinsic photoconductivity. Photoconductivity of undoped $[(C_{16}S)_8Pc]_2Eu$ film was very small (Fig. 6, open circles). We could manage to detect the photocurrent using only in-plane electrode geometry because in this case the dark current was below 1 pA. The redshifted photocurrent peaks corresponding to Soret and *Q* bands were observed in undoped $[(C_{16}S)_8Pc]_2Eu$. In C_{60} -doped $[(C_{16}S)_8Pc]_2Eu$ thin film the strong enhancement of PC was observed, especially in the UV region (Fig. 6, open triangles).

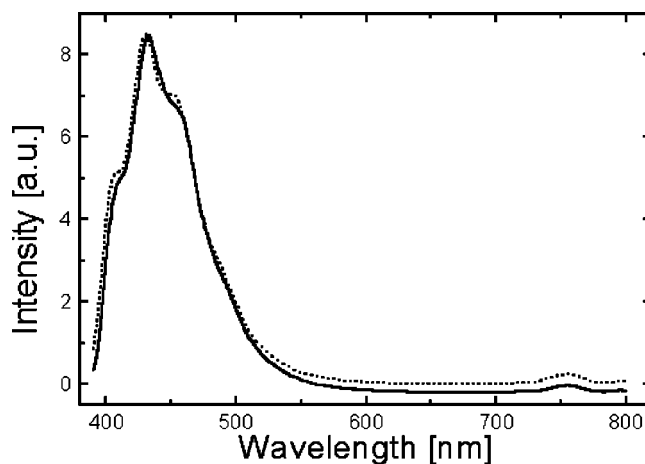


FIG. 5. The PL of $[(C_{16}S)_8Pc]_2M$ infiltrated into disordered opal matrices.

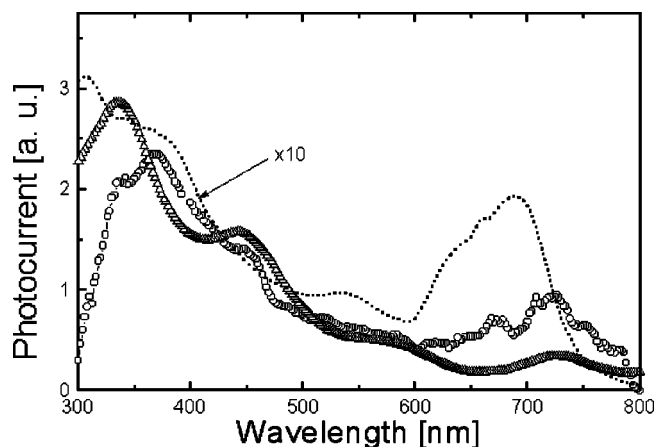


FIG. 6. PC spectra of undoped (open circles) and C_{60} -doped $[(C_{16}S)_8Pc]_2Eu$ (5 mol %), open triangles) and absorption spectra of undoped cast film (dotted line).

The C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ samples show much higher PC. Figure 7 shows the photocurrent action spectra of C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ (20 mol %) for in-plane electrode geometry (open circles) and for sandwich structure (open triangles) and absorption spectra of undoped cast film (dotted line). The sandwich structure was prepared by placing the C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ complex between two patterned indium–tin–oxide (ITO) glass substrates and the sample was heated above solid–liquid crystal phase transition temperature for 1.5 h in Ar atmosphere and then slowly cooled. Again, as in the case of $[(C_{16}S)_8Pc]_2Eu$, the redshifted photocurrent peaks corresponding to Soret and Q bands were observed in C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$. However, the PC was considerably enhanced and new strong peaks at about 260 and 335 nm, corresponding to photoexcitation of C_{60} dopant molecule, were clearly detected. We should note that the parallel to column PC in the sandwich device was measured upon illumination through the ITO electrode, so the photoresponse was cut below 310 nm.

We have observed the effect of negative PC at high voltages (Fig. 8). The corresponding dark current–voltage char-

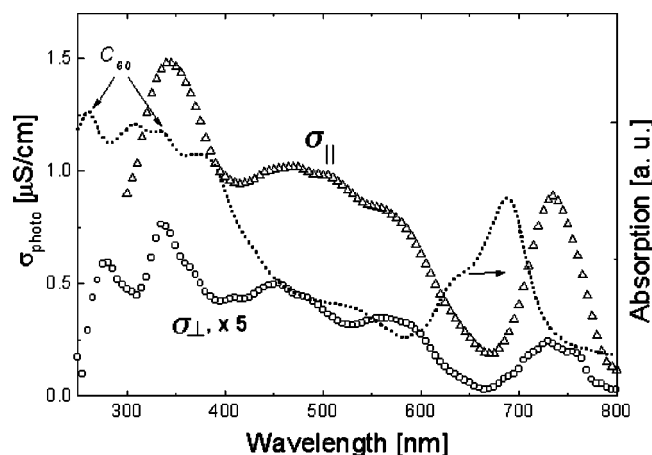


FIG. 7. PC spectra of C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ (20 mol %) (open circles—in-plane electrode geometry, open triangles—sandwich structure) and absorption spectra of undoped cast film (dotted line).

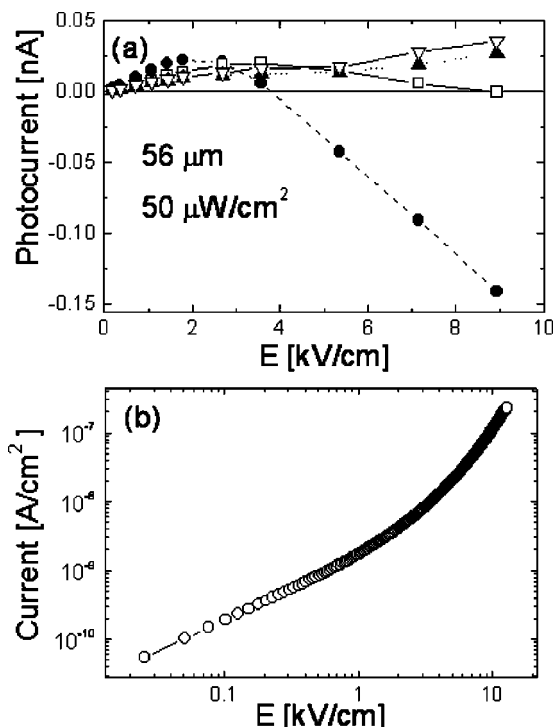


FIG. 8. (a) PC vs electric field dependence of C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ (20 mol %) at different photoexcitation wavelength: 365 nm (solid circle), 470 nm (open square), 560 nm (up triangle), and 735 nm (down triangle). (b) Dark current–voltage characteristic of C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ (20 mol %).

acteristic of C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ (20 mol %) is shown in Fig. 9. The dark current–voltage characteristic is a typical characteristic with those of space charge limited current with the traps in an energy gap.¹⁴ The PC is positive at low voltages in the ohmic regime ($E < 10^3$ V/cm). At higher voltages, when the space charge effects become considerable, the PC decreases for highly absorbing light while it increases for low absorbing light.

IV. DISCUSSION

A. Absorption and photoluminescence

The absorption spectrum of $[(C_{16}S)_8Pc]_2M$ is rather complex. Figure 9 shows the electronic level scheme of the Pc_2Lu dimer. The scheme was drawn according to theoretical calculation by Orti *et al.* for the bis(phthalocyaninato)lutetium complex using the valence effective Hamiltonian non-empirical technique.¹⁶ The close similarity of absorption spectra of all $[(C_{16}S)_8Pc]_2M$ in solution justifies the use of the scheme for all complexes. The optical absorption spectra are only slightly sensitive to the substitution of the central rare-earth metal atom and alkylthio side group.

The so-called Q band in absorption spectrum consists of two sets because of splitting of the phthalocyanine ring energy levels in the dimer and the radical nature of the rare-earth metal phthalocyanine dimer: $Q1$ (710 and 680 nm) due to excitation of electron from the half-occupied HOMO to the next to LUMO excited states and $Q2$ (638 and 572 nm) due to excitation from SOMO to LUMO.

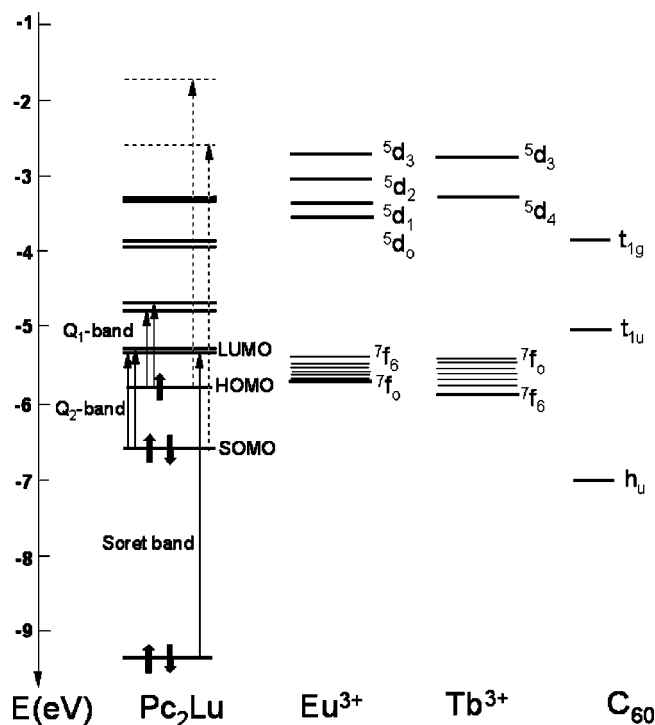


FIG. 9. Electronic energy scheme of Pc_2Lu and Eu^{3+} and Tb^{3+} ions and C_{60} (dashed lines represent less probable processes).

The Soret band in these complexes with a maximum at about 390 nm corresponds to electronic transition from deep level to the LUMO and may be overlapped with electronic transitions from SOMO and HOMO to higher excited states. The small bathochromic shift of Q absorption bands with the increase of atomic number of rare-earth ions has been detected while Soret and high energy absorption peaks were not changed. The effective ionic radii of Eu^{3+} , Tb^{3+} , and Lu^{3+} are 0.947, 0.923, and 0.861 Å, respectively. The similar blueshift upon substitution of the central rare-earth metal atom with Dy, Tm, Yb, and Lu was observed by Markovitsi *et al.*¹⁵ The shoulder at about 430 nm corresponding to transition from the deeper level to the half-occupied HOMO was also observed.¹⁶

PL emission spectra are very weak in comparison with metal free phthalocyanine in solution. Two PL bands were detected simultaneously at around 400–650 and 720–800 nm in chloroform solution upon photoexcitation with a 350 nm laser light. The position and relative intensities of these PL bands correlate well with the corresponding Soret and Q absorption bands, respectively.

The low energy PL band should be discussed in detail. It consists of sharp peak at 740 nm and shoulder at 780 nm. They can be ascribed to emission from next to LUMO high excited states to HOMO and from LUMO to SOMO, respectively, and correspond to relevant absorption peaks of the Q band. This energetic disposition can explain their relative intensity relation. Luminescence from higher energy states is more probable. In the case of excitation to LUMO level the internal energy conversion is highly probable because of the small energy separation between LUMO and HOMO. The internal conversion can be facilitated by strong exchange

coupling induced by the presence of heavy lanthanide ions (heavy-metal atom effect).

Very weak PL intensity suggests that in $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$ the intersystem crossing from the singlet excited state to the triplet excited state should suppress the fluorescence strongly due to the large spin-orbit coupling in the presence of the heavy Eu atom just as in the case of Cu–phthalocyanine.¹⁷ In Cu–phthalocyanine the fluorescence was totally suppressed and the phosphorescence was observed at around 1.12 μm . Indeed, the existence of the large spin-orbit coupling in bis(phthalocyaninato)lanthanide compounds has been shown experimentally.¹⁸ We should mention here that the PL of $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Lu}$ is much higher than that of $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$ since the 4f shell of Lu ion is filled. The phosphorescence of metal phthalocyanine is usually very weak.¹⁹

Emission of the Eu^{3+} ion was not observed compared with other Eu complexes.²⁰ It can be explained if we consider the relative positions of energy levels of the phthalocyanine π -electron system and 4f levels of Eu^{3+} (Fig. 9). Direct excitation of the Eu^{3+} ion has very low probability.¹² Singlet energy transfer is also not significant.²¹ Usually the narrow line emission from the Eu^{3+} ion occurs by energy transfer through the triplet state of the organic unit. However, in the case of $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$ the lowest triplet state is below the resonance level of the Eu^{3+} ion. Therefore, the triplet energy transfer to the Eu^{3+} ion is impossible and the line emission from the Eu^{3+} ion cannot be observed. We should mention here that the PL of the $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Lu}$ complex is higher than that of $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$ since the 4f shell of the Lu^{3+} ion is filled. The strong PL quenching effect in chloroform solution of $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$ was found at increasing concentrations. We should stress here that the study of PL properties of different alkylthio-group-substituted phthalocyanine rare-earth metal sandwich complexes has shown that the effect of rare-earth metal ion is not just a change of the distance between the two phthalocyanine macrocycles but it plays an important role in photoelectronic processes.

B. Photoconductivity

Bis(phthalocyaninato)lanthanide complexes are neutral radicals with the hole delocalized over two phthalocyanine rings due to the interplanar π – π interaction between two phthalocyanine rings. They were considered as a first intrinsic organic semiconductor. However, alkylthio-group-substituted phthalocyanine double decker complexes form quasioone-dimensional columnar structure in the condensed state with a large intercolumnar distance of about 35 Å and show much lower conductivity. It can be considered as a p -type organic insulator. Because of its quasioone-dimensional nature, the electrical transport is dispersive and very sensitive to a number of structural and chemical defects.² Preliminary results on ac conductivity²² have shown the $\omega^{0.8}$ frequency dependence in liquid-crystalline phase similar to other discotic liquid crystals.² The dark current–voltage characteristic indicates the distribution of traps. We believe the model of conductivity by hopping between phthalocyanine rings along columns is also valid for alkylthio-substituted bisphthalocyanine rare-earth metal complexes.²

The identification of the mesophases was carried out by x-ray diffractometry, differential scanning calorimetry (DSC) measurements, and polarization microscopic observations. When the virgin crystal (*K*) was heated at 10 °C/min, it transformed into the discotic hexagonal columnar mesophase D_h , which is stable in a relatively wide temperature range. On further heating, the D_h mesophase cleared into the isotropic liquid (Iso).⁸

X-ray analysis of lanthanide complexes has shown that the average distance between the two phthalocyanine rings is about 3.3–3.5 Å, which leads to relatively large interring overlap integrals in the dimer complexes.⁸

Photoconductivity of undoped $[(C_{16}S)_8Pc]_2M$ compound is very small. Nonradiative relaxation of photoexcitations is dominating. One of the main mechanisms is the intersystem crossing to the lowest triplet state with the following intersystem crossing to the ground state enhanced by strong spin-orbit coupling due to heavy atom effect. The exchange perturbation effect increases the triplet decay significantly. For example, the presence of paramagnetic copper ions in phthalocyanine derivative columnar liquid crystalline phases increased the triplet decay by 2 orders of magnitude.²³

The doping effect of C_{60} in organic materials such as conducting polymers and its enhancing effect on photogeneration of charge carriers have been studied in detail.²⁴ The possibility of *p*- and *n*-type doping of Lutetium bisphthalocyanine has been demonstrated by Maitrot *et al.*²⁵ In the phthalocyanine compounds the photogeneration occurs through excitation of the phthalocyanine molecule to the first excited singlet state, the dissociation of a loosely bound exciton at some dissociation center, or at the interface. The enhancement of photoconductivity in C_{60} -doped $[(C_{16}S)_8Pc]_2M$ is due to the efficient dissociation of excitons by C_{60} . Recently, experimental evidence of charge transfer between C_{60} and phthalocyaninato metal complex was obtained by the photocurrent measurements in Langmuir–Blodgett film of $C_{60}/CuPc(OC_8H_{17})_8$.²⁶

The relative electronic energy levels of the C_{60} -doped complex is shown in Fig. 9. The remarkable peculiarity of this phthalocyanine compound is the strong effect of C_{60} doping on the PC spectrum within the Soret band but lower in the Q band. This feature can be explained by considering the electronic scheme of the C_{60} -doped $[(C_{16}S)_8Pc]_2M$ complex. The LUMO of the C_{60} molecule lies a little higher than the LUMO of $[(C_{16}S)_8Pc]_2M$ but lower than the next excited states of $[(C_{16}S)_8Pc]_2M$. The Q band in the absorption of phthalocyanine complex corresponds to photoexcitation of the electron from SOMO to LUMO and from HOMO to next LUMO. According to the relative position of LUMO of C_{60} , C_{60} can only partially affect the photoresponse of the Q band.

The origin of photoresponse with a maximum at 500–600 nm is not clear. There are two possibilities. One is the partially allowed h_u-t_{1u} molecular transition of C_{60} . The h_u-t_{1u} molecular transition of C_{60} is forbidden. However, it may be partially allowed because of the Jahn–Teller effect and electron–phonon coupling due to the sensitivity of C_{60} and its anion to environment.⁹ The extra charge on the C_{60} ball can also cause the enhancement of vibrational modes.

The second possibility is the triplet–triplet excitation.²⁷

On the contrary, the higher excited states dissociate more easily. The new strong peak at about 340 nm shows that the optical excitation at photon energy corresponding to the allowed transition (h_u-t_{1g}) in C_{60} causes the remarkable PC enhancement. In this case, a hole will be transferred from C_{60} to $[(C_{16}S)_8Pc]_2M$. An excited electron on C_{60} relaxes into the negative polaron on C_{60} .

One of the most noticeable features of PC spectra is the redshift of PC peaks corresponding to intense Soret and *B* absorption bands. The most probable origin of this effect is the high rate of surface recombination. The observation of the effect of the negative PC strongly supports this explanation.

The effect of negative PC was observed in a number of materials and its origin may be very different and complicated. One of the possible mechanisms of the negative PC in C_{60} -doped $[(C_{16}S)_8Pc]_2M$ is the following. Upon photoexcitation the photoinduced electron transfer occurs from $[(C_{16}S)_8Pc]_2M$ to C_{60} . It is well known that such a forward photoinduced electron transfer is very fast²⁸ and the electron relaxes to a polaron level. Therefore back electron transfer is delayed. At the same time the free hole remaining on phthalocyanine complex is trapped by some level in energy gap below the Fermi level. Then this trapped hole can recombine with free electron decreasing the steady state current below the dark current. A similar process may occur with holes. In this case, upon photoexcitation the photoinduced hole transfer occurs from C_{60} to $[(C_{16}S)_8Pc]_2M$ and the electron on C_{60} relaxes to the polaron level. Kirova proposed²⁹ that the capture cross section for electrons and holes may not be different for such large size organic molecules as C_{60} . This may occur for the charge transfer complexes of very large molecules such as the substituted phthalocyanine rare-earth metal compound and C_{60} . However, this model can explain the negative PC only at negative voltages applied to the illuminated ITO electrode, while the effect of negative PC was observed at both-polarities of applied voltages, although it was about two times smaller for negative voltages. Moreover, the electric field and wavelength dependence of negative photoconductivity of C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ definitely supports the space charge effect formed by efficient charge transfer between C_{60} and $[(C_{16}S)_8Pc]_2Lu$. The photoconductivity is positive at low voltages in the ohmic regime. At relatively high electric fields, buildup of space charge occur at the illuminated ITO electrode–organic film interface. The concentration of 20 mol % corresponds to one C_{60} molecule for five dimer complexes. Hence the significant space charge and strong Coulomb field can be built at the interface. The corresponding dark current–voltage characteristic of undoped and C_{60} -doped $[(C_{16}S)_8Pc]_2Lu$ shows the space charge limited current with the traps in energy gap; it is linear at low field ($<10^{-3}$ V/cm) but became supralinear $I-V^n$ with $n \sim 3$ at high electric field.

The peculiarity of the substituted phthalocyanine rare-earth metal complexes is that the dark conductivity and mobility increases with temperature while most discotic materials have shown a decrease in mobility. For example, the mobility of $[(C_{12}O)_8Pc]_2Lu$ increases from $\sim 2 \times 10^{-6}$ in

crystalline state to $1.7 \times 10^{-5} \text{ m}^2/(\text{Vs})$ in discotic phase because of the structural reorganization from tilted to horizontally stacked, hexagonally packed columnar liquid crystalline phase and the equivalence of phthalocyanine macrocycles within the column.¹⁸ A study of photoconductivity in the liquid crystalline phase will be published elsewhere.²²

The electronic energy scheme of substituted phthalocyanine rare-earth metal sandwich complexes is very complicated and further study is needed to clarify the electronic processes in the compounds. The x-ray structural analysis has shown that two Pc macrocycles are not coplanar but concave with one macrocycle slightly more distorted than the other because of the radical nature of this complex.³⁰ The rare-earth metal ion permits charge transfer between two phthalocyanine ligands in the double decker molecule.³ In this article we did not discuss the intramolecular charge transfer from SOMO to the half-occupied HOMO polarized along the columnar axis in the near infrared region.^{8,15}

V. CONCLUSIONS

A comparative study of optical properties of alkylthio-group-substituted phthalocyanine rare-earth metal sandwich complexes such as $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$, $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Tb}$, and $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Lu}$ was presented. PL and PC of the $[(\text{C}_n\text{S})_8\text{Pc}]_2\text{M}$ complex is very weak.

Two PL bands were observed at around 400–650 and 720–800 nm in chloroform solution corresponding to the Soret and Q bands in the absorption spectra, respectively. However, the emission from the Eu^{3+} ion was not observed in contrast to other Eu complexes. The strong PL quenching effect in chloroform solution of $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Eu}$ was observed at increased concentrations due to the aggregation effect.

We have shown the C_{60} doping effect in thin film of the columnar $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{M}$ complex. The significant enhancement of photoconductivity after C_{60} doping is reported. In the case of the $[(\text{C}_{16}\text{S})_8\text{Pc}]_2\text{Lu}$ complex the enhancement of PC measured parallel to columns by more than 2 orders of magnitude, was observed. The PC is positive at low electric field in the ohmic regime while it becomes negative at high electric field upon photoexcitation with strongly absorbed light. The negative PC is attributed to space-charge effects.

The results of PC on C_{60} doped substituted phthalocyanine rare-earth metal complexes have shown that the materials of self-assembled columnar stacks with large intercolumnar spaces could be significantly photosensitized and present a very promising class of photoactive material.

ACKNOWLEDGMENTS

This work was financially supported by the Research for the Future Program of the Japan Society for the Promotion of

Science (Project No. JSPS-RFTF96P00206). This work was also partially supported by a Grant-in-Aid for COE Research (Grant No. 10CE2003) by the Ministry of Education, Science, Sports and Culture of Japan. The authors would like to thank H. Kajii for help in sample preparation.

- ¹D. Adam, P. Schuhmacher, J. Simmerer, L. Häussling, K. Siemensmeyer, K. H. Etzbach, H. Ringsdorf, and D. Haarer, *Nature (London)* **371**, 141 (1994).
- ²N. Boden, R. J. Bushby, J. Clements, B. Movaghar, K. J. Donovan, and T. Kreouzis, *Phys. Rev. B* **52**, 13274 (1995).
- ³A. M. van de Craats, J. M. Warman, H. Hasebe, R. Naito, and K. Ohta, *J. Phys. Chem.* **101**, 9224 (1997).
- ⁴K. Ohta, L. Jacquemin, C. Sirlin, L. Bsio, and J. Simon, *New J. Chem.* **12**, 751 (1988).
- ⁵C. Videlot, D. Fichou, and F. Garnier, *Synth. Met.* **102**, 1052 (1999).
- ⁶K. Petritsch, R. H. Friend, A. Lux, G. Rozenberg, S. C. Moratti, and A. B. Holmes, *Synth. Met.* **102**, 1776 (1999).
- ⁷Z. Belarbi, M. Maitrot, K. Ohta, J. Simon, J.-J. Andre, and P. Petit, *Chem. Phys. Lett.* **143**, 400 (1988).
- ⁸K. Ban, K. Nishizawa, and K. Ohta (unpublished).
- ⁹K. Yoshino, S. Lee, A. Fujii, H. Nakayama, W. Schneider, A. Naka, and M. Ishikawa, *Adv. Mater.* **11**, 1382 (1999).
- ¹⁰L. Edwards and M. Gouterman, *J. Mol. Spectrosc.* **33**, 292 (1970).
- ¹¹B. H. Schechtman and W. E. Spicer, *J. Mol. Spectrosc.* **33**, 28 (1970).
- ¹²M. H. V. Werts, J. W. Hofstra, F. A. J. Geurts, and J. W. Verhoeven, *Chem. Phys. Lett.* **276**, 196 (1997).
- ¹³M. Fujiki, H. Tabei, and T. Kurihara, *J. Phys. Chem.* **92**, 1281 (1988).
- ¹⁴M. A. Lampert and P. Mark, *Current Injection in Solids* (Academic, New York, 1970).
- ¹⁵D. Markovitsi, T.-H. Tran-thi, R. Even, and J. Simon, *Chem. Phys. Lett.* **137**, 107 (1987).
- ¹⁶E. Orti, J. L. Bredas, and C. Clarisse, *J. Chem. Phys.* **92**, 1228 (1990).
- ¹⁷P. G. Schouten, J. M. Warman, M. P. de Haas, M. A. Fox, and H.-L. Pan, *Nature (London)* **353**, 736 (1991).
- ¹⁸K. L. Trojan, W. E. Hatfield, K. D. Kepler, and M. L. Kirk, *J. Appl. Phys.* **69**, 6007 (1991).
- ¹⁹K. Yoshino, K. Kaneto, J. Kyokane, and Y. Inuishi, *J. Phys. Soc. Jpn.* **31**, 1594 (1971).
- ²⁰See, for example, Y. Ohmori, H. Ueta, Y. Kurosaka, Y. Hamada, and K. Yoshino, *Jpn. J. Appl. Phys., Part 2* **37**, L798 (1998).
- ²¹G. A. Crosby, R. E. Whan, and R. M. Alire, *J. Chem. Phys.* **34**, 743 (1961).
- ²²K. Yoshino, S. Lee, T. Sonoda, K. Nakayama, M. Ozaki, K. Ban, and K. Ohta (unpublished).
- ²³D. Markovitsi, I. Lecuyer, and J. Simon, *Chem. Phys. Lett.* **167**, 467 (1990).
- ²⁴S. Morita, A. A. Zakhidov, and K. Yoshino, *Solid State Commun.* **82**, 249 (1992).
- ²⁵M. Maitrot, G. Guillaud, B. Boudjema, J.-J. Andre, H. Strzelecka, J. Simon, and R. Even, *Chem. Phys. Lett.* **133**, 59 (1987).
- ²⁶D. Tanaka, M. Rikukawa, K. Sanui, and N. Ogata, *Synth. Met.* **102**, 1492 (1999).
- ²⁷D. Markovitsi, I. Lecuyer, and J. Simon, *J. Phys. Chem.* **95**, 3620 (1991).
- ²⁸N. S. Sariciftci, L. Smilowitz, A. J. Heeger, and F. Wudl, *Science* **258**, 1474 (1992).
- ²⁹N. Kirova and S. Brazovskii, *Synth. Met.* **76**, 229 (1996).
- ³⁰A. de Cian, M. Monsavi, J. Fischer, and R. Weiss, *Inorg. Chem.* **24**, 3162 (1985).