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Osaka University

# **Doctoral Dissertation**

## **Studies on Novel Catalysts Based on Zirconium-Tin Oxide for Direct Decomposition of Nitrous Oxide**

(ジルコニウム-スズ複合酸化物を用いた  
新規な亜酸化窒素直接分解触媒に関する研究)

**HU DONGRUN**

**January 2024**

Department of Applied Chemistry  
Graduate School of Engineering  
Osaka University

## ***Preface***

The study of this doctoral thesis has been conducted under the direct supervision of Professor Dr. Nobuhito Imanaka at Department of Applied Chemistry, Graduate School of Engineering, Osaka University.

The object of the present thesis is to develop novel catalysts based on novel materials of zirconium-tin oxide for direct decomposition of nitrous oxide.

The author desires that the findings and acknowledge obtained in this study will contribute to the further design and development of catalysts for nitrous oxide decomposition and the materials would be widespread application in the future.

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January 2024

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## ***General Introduction***

Nitrous oxide ( $\text{N}_2\text{O}$ ) is a well-known environmental pollutant that causes the global warming, and its greenhouse effect is ca. 300 times higher than that of carbon dioxide [1,2]. Also,  $\text{N}_2\text{O}$  is a stratospheric ozone-depleting gas, similar to chlorofluorocarbons [3]. Since  $\text{N}_2\text{O}$  is emitted by manufacturing processes, such as production of adipic acid and nitric acid, recent development of industry causes the increase of  $\text{N}_2\text{O}$  concentration in the atmosphere [4,5]. In order to prevent global warming as well as ozone layer depletion, it is important to decompose  $\text{N}_2\text{O}$  from industrial exhaust gases.

Several methods for  $\text{N}_2\text{O}$  decomposition have been reported, including selective catalytic reduction (SCR) [6-8], thermal decomposition [9], and direct catalytic decomposition [10-35]. In the SCR method,  $\text{N}_2\text{O}$  is reduced by a reducing agent, such as  $\text{NH}_3$ , into nontoxic  $\text{N}_2$  and  $\text{H}_2\text{O}$  over SCR catalysts. However, a continuous supply of the reducing additive is required. Since  $\text{NH}_3$  has toxicity, corrosiveness, and flammability, the control system for safety is indispensable. For thermal decomposition,  $\text{N}_2\text{O}$  is decomposed into nontoxic  $\text{N}_2$  and  $\text{O}_2$  by applying elevated temperatures over 800 °C. Although the decomposition can proceed without using reducing agents, the high temperatures cause heat deterioration in the system and require significant energy consumption. In contrast, direct catalytic decomposition of  $\text{N}_2\text{O}$  into  $\text{N}_2$  and  $\text{O}_2$  is considered to be a promising method, because there is no need for reducing agents and the required temperature is moderate compared to the case of thermal decomposition.

According to previous studies, various catalysts for direct N<sub>2</sub>O decomposition have been studied. It is known that Rh based catalysts are effective to decompose N<sub>2</sub>O [10-17], as listed in **Table G-1**. In the Rh based systems, high catalytic activities have been reported for the catalysts having high specific surface areas, such as 2 wt% Rh/USY (complete decomposition: 250 °C, specific surface area: 695 m<sup>2</sup>·g<sup>-1</sup>). However, the morphology of the support material is strictly controlled, because a decrease in the specific surface area generally induces the deactivation owing to a decrease in the number of active sites, caused by a Rh aggregation. The reported catalysts having low specific surface areas show low activities even with noble metals, such as 1 wt% Rh/ZrO<sub>2</sub> (complete decomposition: 550 °C, specific surface area: 79 m<sup>2</sup>·g<sup>-1</sup>). For noble-metal-free catalysts, spinel-type oxides [18-21], perovskite-type oxides [22-25], hexaaluminates [26-29], and zeolites [30-33] have been reported,

**Table G-1.** Reported Rh based catalysts for N<sub>2</sub>O decomposition

Catalyst	Complete decomposition temperature / °C	Specific surface area / m <sup>2</sup> ·g <sup>-1</sup>	Reaction condition	Ref.
2 wt% Rh/USY <sup>a</sup>	250	695	0.095% N <sub>2</sub> O/He, 60000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[10]
2 wt% Rh/Al <sub>2</sub> O <sub>3</sub>	300	168	0.095% N <sub>2</sub> O/He, 60000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[10]
2 wt% Rh/Mg/Al	400	289	0.1% N <sub>2</sub> O/N <sub>2</sub> , 480000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[11]
1 wt% Rh/Ag/Al <sub>2</sub> O <sub>3</sub>	400	138	0.1% N <sub>2</sub> O/He, 300000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[12]
1 wt% Rh/KIT-6 <sup>b</sup>	420	602	0.17% N <sub>2</sub> O/He, 120000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[13]
1wt%Rh/Al <sub>2</sub> O <sub>3</sub>	550	110	1% N <sub>2</sub> O/He, 1800 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[14]
1wt%Rh/ZrO <sub>2</sub>	550	79	1% N <sub>2</sub> O/He, 1800 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[14]

<sup>a</sup> USY: Ultrastable Y zeolite, <sup>b</sup> KIT-6: Korea Advanced Institute of Science and Technology No. 6

**Table G-2.** Typical noble-metal-free catalysts for N<sub>2</sub>O decomposition

Catalyst	Complete decomposition temperature / °C	Specific surface area / m <sup>2</sup> ·g <sup>-1</sup>	Reaction condition	Ref.
Co <sub>3</sub> O <sub>4</sub> (Spinel-type oxide)	500	29	0.1% N <sub>2</sub> O/He, 20000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[21]
LaCoO <sub>3</sub> (Perovskite-type oxide)	550	8	0.5% N <sub>2</sub> O/He, 120000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[22]
LaFeAl <sub>11</sub> O <sub>19</sub> (Hexaaluminate)	825	11	0.15% N <sub>2</sub> O/He, 60000 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[29]
Co-clinoptilolite (Zeolite)	550	27	0.04% N <sub>2</sub> O/N <sub>2</sub> , 4500 mL·g <sup>-1</sup> ·h <sup>-1</sup>	[30]

where typical catalysts are tabulated in **Table G-2**. Although their activities are lower than those of the Rh based catalysts having high specific surface areas, the spinel-type Co<sub>3</sub>O<sub>4</sub> catalyst has effective N<sub>2</sub>O decomposition ability compared to the other metal oxides.

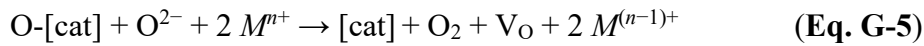
The reaction mechanism for N<sub>2</sub>O decomposition has been suggested as follows [34,35]: N<sub>2</sub>O molecules are adsorbed on the catalytically active sites ([cat]) (**Eq. G-1**), decomposition of adsorbed N<sub>2</sub>O into gaseous N<sub>2</sub> and adsorbed oxygen (**Eq. G-2**), and reaction of adsorbed oxygen with gaseous N<sub>2</sub>O (**Eq. G-3**) or combination of two adsorbed oxygen species (**Eq. G-4**).



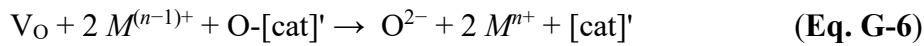
In the above reaction mechanism, the regeneration of the catalytically active sites via removing adsorbed oxygen (**Eqs. G-3 and G-4**) is known to be the rate-determining step [34,35]. One effective



method to facilitate the removal of residual oxygen species is the introduction of promoters with the ability to supply oxygen species from inside the lattice. The proposed reaction is as follows: reaction of the adsorbed oxygen species and an oxide ion in the promoter to form gaseous O<sub>2</sub> with the generation of oxygen vacancies (V<sub>O</sub>) and the reduction of M<sup>n+</sup> (cation in the promoter):



Oxygen vacancies and electrons migrate through the lattice, and they are compensated by oxygen species adsorbed on another active site ([cat]')



Based on the mechanism, the oxygen supply ability of the promoter should be related to the redox properties and the oxide ion conducting property.

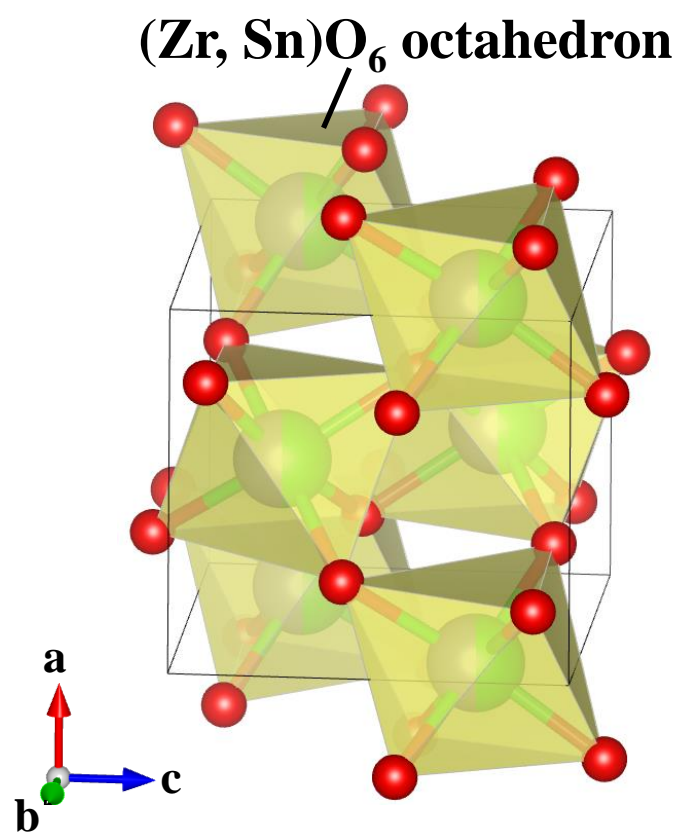
In the Imanaka laboratory, to which I belong, single-phase zirconium-tin oxide (ZrSnO<sub>4</sub>) with an α-PbO<sub>2</sub>-type structure was firstly synthesized, and its crystal structure was demonstrated with Rietveld analysis (**Figure G-1**) [36]. This material may have redox properties because of the valence change between Sn<sup>4+</sup> and Sn<sup>2+</sup>. In addition, the coexistence of Sn<sup>2+</sup> can generate oxygen vacancies for oxide ion migration. Therefore, ZrSnO<sub>4</sub> is expected to have the oxygen supply ability. In this doctoral dissertation, ZrSnO<sub>4</sub> was focused to develop novel catalysts for direct decomposition of N<sub>2</sub>O.

This thesis consists of the following three chapters.

**Chapter 1** aims to develop novel catalysts with high catalytic activity by selecting ZrSnO<sub>4</sub> as a promoter. By combining the ZrSnO<sub>4</sub> promoter and a Rh activator having N<sub>2</sub>O decomposition ability, Rh/ZrSnO<sub>4</sub> catalysts were prepared, and their N<sub>2</sub>O direct decomposition activities were investigated.

In **Chapter 2**, in order to develop the catalysts without using noble metals,  $\text{Co}_3\text{O}_4$  was selected as an activator instead of Rh. In addition, to obtain the high catalytic activity,  $\text{Fe}^{2+/3+}$  ions were introduced into  $\text{ZrSnO}_4$ , because the valence changes of  $\text{Fe}^{2+/3+}$  may improve the redox properties and the replacement of  $\text{Sn}^{4+}$  for lower-valent  $\text{Fe}^{2+/3+}$  can generate oxygen vacancies for smooth oxide ion migration. Therefore,  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts were prepared, and the catalytic activities for  $\text{N}_2\text{O}$  direct decomposition were investigated.

**Chapter 3** deals with the catalysts with durability against coexisting gases of oxygen gas and water vapor. To this end,  $\text{Co}^{2+/3+}$  doping was carried out into the  $\text{ZrSnO}_4$  lattice, instead of  $\text{Co}^{2+/3+}$  loading as  $\text{Co}_3\text{O}_4$ . The introduction of  $\text{Co}^{2+/3+}$  was expected to improve the redox properties due to two types of valence states ( $\text{Co}^{2+}$  and  $\text{Co}^{3+}$ ) and to generate oxygen vacancies caused by the lower-valence states of  $\text{Co}^{2+/3+}$  than that of  $\text{Sn}^{4+}$ . Therefore,  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  solids were prepared, and their  $\text{N}_2\text{O}$  direct decomposition activities were investigated.



**Figure G-1.** Crystal structure of ZrSnO<sub>4</sub> [36].

## ***List of Publications***

1. Direct Decomposition of Nitrous Oxide by Rhodium Supported on  $\text{ZrSnO}_4$

**Dong-Run Hu**, Naoyoshi Nunotani, and Nobuhito Imanaka\*

Functional Materials Letters, **2023**, *16 (3&4)*, 2350016:1–4.

2. Direct Decomposition of  $\text{N}_2\text{O}$  over  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  Catalysts

**Dong-Run Hu**, Naoyoshi Nunotani, and Nobuhito Imanaka\*

Functional Materials Letters, **2023**, *16 (6)*, 2350023:1–4.

3. Novel Cobalt-doped  $\text{ZrSnO}_4$  Catalysts for Direct Nitrous Oxide Decomposition

**Dong-Run Hu**, Naoyoshi Nunotani, and Nobuhito Imanaka\*

International Journal of Applied Ceramic Technology, **2023**, *20 (2)*, 966–971.

## ***Chapter 1***

### **Novel Catalysts of Rh/ZrSnO<sub>4</sub> for Direct Decomposition of N<sub>2</sub>O**

#### **1.1. Introduction**

As described in the General Introduction section, the promoter which can supply oxygen species from inside the lattice may facilitate N<sub>2</sub>O decomposition. In addition, ZrSnO<sub>4</sub> is expected to have oxygen supply ability due to its redox properties through the valence change between Sn<sup>4+</sup> and Sn<sup>2+</sup> and oxygen vacancies for oxide ion migration formed by the coexistence of Sn<sup>2+</sup>.

In Chapter 1, in order to develop novel catalysts with high catalytic activity, the ZrSnO<sub>4</sub> solid was selected as a promoter, and a Rh activator having N<sub>2</sub>O decomposition ability was loaded onto ZrSnO<sub>4</sub>. Therefore, Rh/ZrSnO<sub>4</sub> catalysts were prepared, and their catalytic activities for N<sub>2</sub>O decomposition were investigated.

#### **1.2. Experimental Procedure**

The ZrSnO<sub>4</sub> solid was prepared using a co-precipitation method as follows. Powders of zirconium (IV) oxynitrate dihydrate (ZrO(NO<sub>3</sub>)<sub>2</sub>·2H<sub>2</sub>O, Kishida Chemical, ≥ 99.0%) (0.4864 g) and tin (II) oxalate (SnC<sub>2</sub>O<sub>4</sub>, Sigma-Aldrich, 99%) (0.2891 g) were weighed, where a Zr:Sn ratio was 1.3:1. They were added into deionized water, and then nitric acid solution (Kishida Chemical, 60%) (10 mL) was poured into the mixture to obtain a homogenous solution, followed by stirring at room

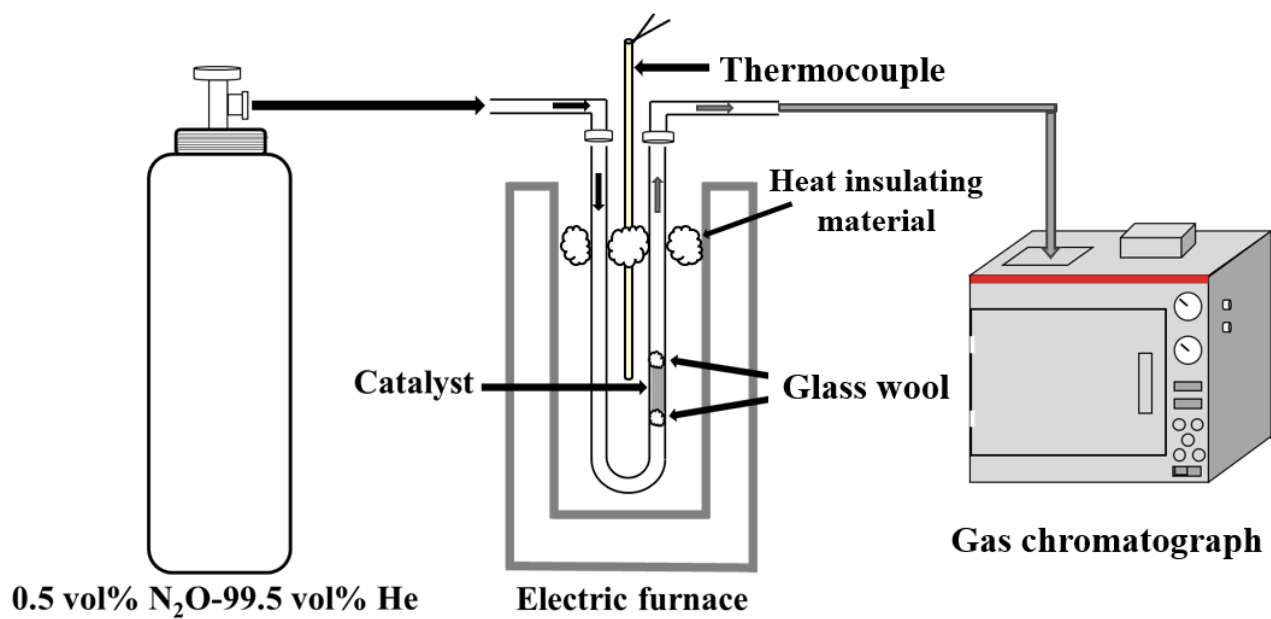
temperature for 30 min. Subsequently, the pH was adjusted to 11 by dropwise addition of an aqueous ammonia solution (Kishida Chemical, 28%). After stirring for 6 h at room temperature, the resulting precipitate was collected by filtration, washed with deionized water, and then dried at 80 °C for 6 h. Finally, the powder was calcined at 600 °C for 1 h under air flow (10 mL·min<sup>-1</sup>).

Rh supported on catalysts were prepared via impregnating the ZrSnO<sub>4</sub> powders (0.3000 g) with 0.01 mol·L<sup>-1</sup> RhCl<sub>3</sub> aqueous solution (prepared by dissolving rhodium (III) chloride hydrate (RhCl<sub>3</sub>·*n*H<sub>2</sub>O, Kishida Chemical) in deionized water). After the evaporation of the solvent at 180 °C using a hot stirrer, the residue was ground by using an agate mortar, followed by the calcination at 500 °C for 2 h in the air to obtain *x* wt% Rh/ZrSnO<sub>4</sub> (*x* = 0.5, 1.0, 1.5, and 2.0). For comparison purpose, 1.0 wt% Rh/SnO<sub>2</sub>, 1.0 wt% Rh/ZrO<sub>2</sub>, and 1.0 wt% Rh/Al<sub>2</sub>O<sub>3</sub> were prepared by a similar impregnation method using tin (IV) oxide (SnO<sub>2</sub>, Kojundo Chemical Laboratory, ≥ 99.9%), zirconium (IV) oxide (ZrO<sub>2</sub>, Kanto Chemical, ≥ 99.9%) and aluminum (III) oxide (Al<sub>2</sub>O<sub>3</sub>, Kojundo Chemical Laboratory, ≥ 99.99%).

The crystal structures were identified by using X-ray powder diffraction (XRD; SmartLab, Rigaku) with Cu K $\alpha$  radiation (40 kV and 30 mA) in the 2 $\theta$  range between 10 and 70 deg. The XRD data were analyzed by whole powder pattern fitting using PDXL software (Rigaku) to calculate the crystallite size. The Rh loading amount and sample compositions were determined by X-ray fluorescence spectroscopy (XRF; Supermini200, Rigaku) analysis. The Brunauer–Emmett–Teller (BET) specific surface area was determined at -196 °C in a nitrogen environment (Micromeritics TriStar 3000 analyzer, Shimadzu). The Rh dispersion was determined using a CO-pulse method at

50 °C with 10 vol% CO - 90 vol% He (0.03 mL) over 0.05 g of the catalyst, under the suppose that one CO molecule adsorbs to one surface Rh atom (BELCAT-B, MicrotracBEL). After preheating at 200 °C for 1 h under a helium flow (50 mL·min<sup>-1</sup>), hydrogen temperature-programmed reduction (TPR) measurement was carried out under a reducing gas flow (5 vol% H<sub>2</sub> - 95 vol% Ar at 50 mL·min<sup>-1</sup>) at a heating rate of 5 °C·min<sup>-1</sup> over 0.05 g of the sample (BELCAT-B, MicrotracBEL). The valence state of surface Rh species was investigated by X-ray photoelectron spectroscopy (XPS; PHI5000 VersaProbe II, ULVAC-PHI) with Al K $\alpha$  radiation, and the effect of charging on the binding energies was refined by the C 1s peak (284.6 eV), and the spectrum was fitted using a Shirley background and Gaussian–Lorentzian line shapes.

The catalytic activity for N<sub>2</sub>O direct decomposition was investigated using a conventional fixed-bed flow reactor with a silica-glass tube (10 mm diameter). **Figure 1-1** illustrates a schematic view of the fixed-bed flow reactor for the N<sub>2</sub>O direct decomposition. The reactant gas of 0.5 vol% N<sub>2</sub>O - 99.5 vol% He was fed at 60 mL·min<sup>-1</sup> over 0.2 g of the catalyst with a catalyst weight hourly space velocity of 18,000 mL·g<sup>-1</sup>·h<sup>-1</sup>. Before the catalytic activity tests, each catalyst was preheated at 200 °C for 1 h under a helium flow (50 mL·min<sup>-1</sup>). After the catalytic reaction, the N<sub>2</sub>O conversion was determined using a gas chromatograph equipped with a thermal conductivity detector (GC-8AIT, Shimadzu).



**Figure 1-1.** A schematic view of the fixed-bed flow reactor for  $\text{N}_2\text{O}$  direct decomposition.



### 1.3. Results and Discussion

Figure 1-2 shows the XRD patterns of the  $x$  wt% Rh/ZrSnO<sub>4</sub> ( $x = 0.5, 1.0, 1.5, 2.0$ ) catalysts with the data of ZrSnO<sub>4</sub>. The ZrSnO<sub>4</sub> solid exhibited the single-phase of the  $\alpha$ -PbO<sub>2</sub>-type structure. For the  $x$  wt% Rh/ZrSnO<sub>4</sub> catalysts, peaks were identified as only the  $\alpha$ -PbO<sub>2</sub>-type phase, and no peak related to the Rh species was detected. This was probably due to the low loading amount of Rh.

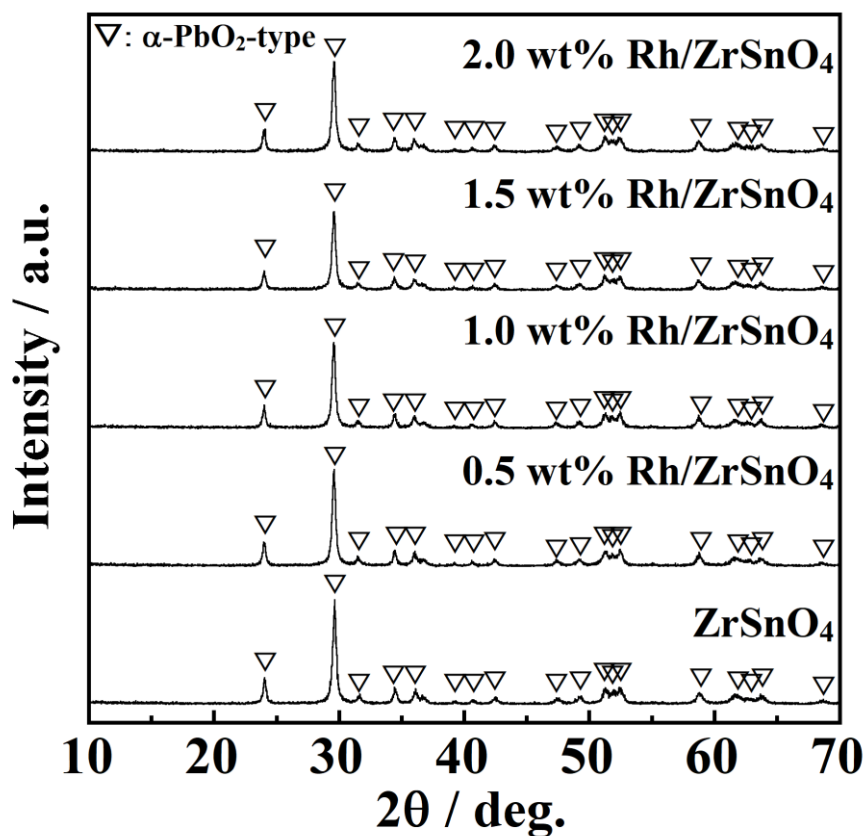


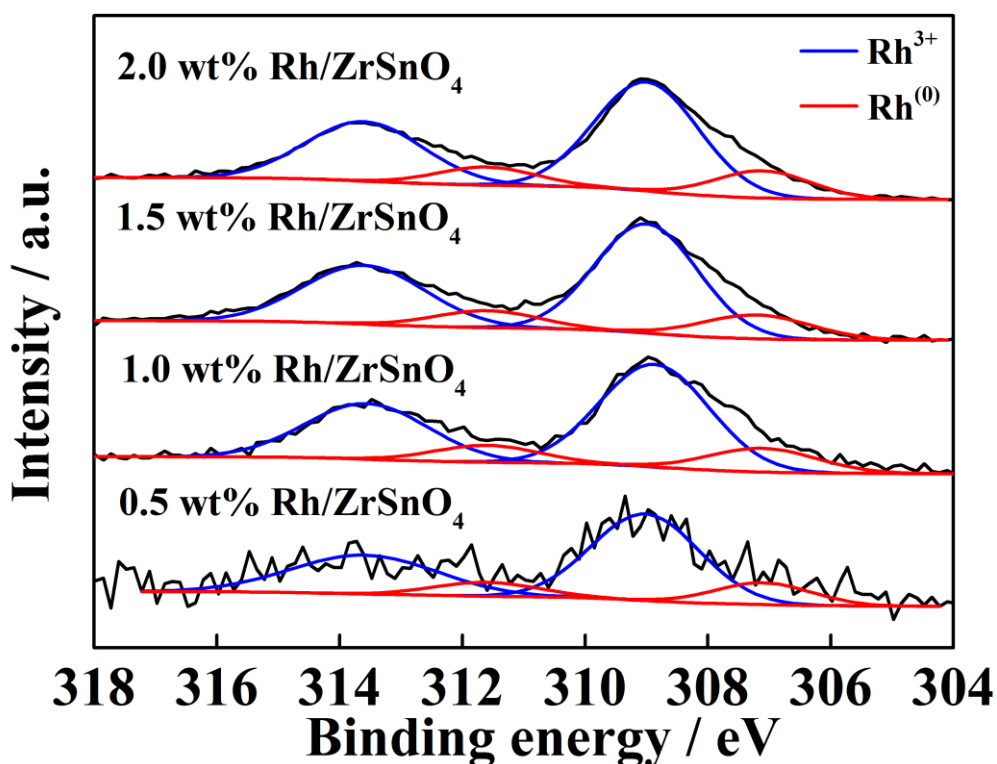
Figure 1-2. XRD patterns of  $x$  wt% Rh/ZrSnO<sub>4</sub> and ZrSnO<sub>4</sub>.

The measured composition, specific surface area, crystallite size of ZrSnO<sub>4</sub>, and Rh dispersion of the  $x$  wt% Rh/ZrSnO<sub>4</sub> catalysts are summarized in **Table 1-1**. The measured compositions were in good agreement with the feed values within the experimental error; that is, Rh was successfully loaded onto the ZrSnO<sub>4</sub> solid. For the specific surface area, all the samples showed similar values of ca. 13 m<sup>2</sup>·g<sup>-1</sup>. Furthermore, similar crystallite sizes of ZrSnO<sub>4</sub> (ca. 32 nm) were obtained. These results might be explained by the fact that ZrSnO<sub>4</sub> was prepared by the same method before loading Rh. In contrast, the Rh dispersion decreased monotonously with increasing the Rh loading amount ( $x$ ), suggesting the Rh aggregation on the ZrSnO<sub>4</sub> surface.

**Table 1-1.** Measured composition, specific surface area, crystallite size of ZrSnO<sub>4</sub>, and Rh dispersion of the  $x$  wt% Rh/ZrSnO<sub>4</sub> catalysts

Feed composition	Measured composition	Specific surface area / m <sup>2</sup> ·g <sup>-1</sup>	Crystallite size of ZrSnO <sub>4</sub> / nm	Rh dispersion / %
ZrSnO <sub>4</sub>	Zr <sub>1.07</sub> Sn <sub>0.93</sub> O <sub>4</sub>	14.0	33	-
0.5 wt% Rh/ZrSnO <sub>4</sub>	0.52 wt% Rh/Zr <sub>1.07</sub> Sn <sub>0.93</sub> O <sub>4</sub>	13.0	32	53.2
1.0 wt% Rh/ZrSnO <sub>4</sub>	0.96 wt% Rh/Zr <sub>1.08</sub> Sn <sub>0.92</sub> O <sub>4</sub>	13.6	33	32.8
1.5 wt% Rh/ZrSnO <sub>4</sub>	1.56 wt% Rh/Zr <sub>1.03</sub> Sn <sub>0.97</sub> O <sub>4</sub>	12.0	31	24.8
2.0 wt% Rh/ZrSnO <sub>4</sub>	2.00 wt% Rh/Zr <sub>1.06</sub> Sn <sub>0.94</sub> O <sub>4</sub>	12.0	30	21.3

In order to investigate the oxidation states of the surface Rh species, XPS measurements were performed for the  $x$  wt% Rh/ZrSnO<sub>4</sub> catalysts. **Figure 1-3** presents the XPS results of the Rh 3d core levels, where the fitting parameters are tabulated in **Table 1-2**. For all the samples, the Rh<sub>3/2</sub> and Rh<sub>5/2</sub> peaks could be divided into Rh<sup>3+</sup> and Rh<sup>(0)</sup> peaks. The ratios of Rh<sup>3+</sup>/(Rh<sup>3+</sup>+Rh<sup>(0)</sup>) were calculated by the peak areas related to Rh<sup>3+</sup> and Rh<sup>(0)</sup> in the XPS spectra, and the results are listed in **Table 1-3**. The Rh<sup>3+</sup>/(Rh<sup>3+</sup>+Rh<sup>(0)</sup>) ratios for all the samples were almost the same values (ca. 81%), irrespective of the Rh loading amount.



**Figure 1-3.** XPS spectra of the Rh 3d core levels of the  $x$  wt% Rh/ZrSnO<sub>4</sub> catalysts.

**Table 1-2.** Fitting parameters of the Rh 3d XPS spectra for  $x$  wt% Rh/ZrSnO<sub>4</sub>

Catalyst	$\chi^2$	Peak	Position / eV	FWHM / eV	Area / %
0.5 wt% Rh/ZrSnO <sub>4</sub>	0.294	Rh <sup>3+</sup> 3d <sub>5/2</sub>	309.0	2.1	15.0
		Rh <sup>3+</sup> 3d <sub>3/2</sub>	313.6	2.8	9.2
		Rh <sup>(0)</sup> 3d <sub>5/2</sub>	307.1	1.9	3.4
		Rh <sup>(0)</sup> 3d <sub>3/2</sub>	311.6	2.0	2.4
1.0 wt% Rh/ZrSnO <sub>4</sub>	0.037	Rh <sup>3+</sup> 3d <sub>5/2</sub>	308.9	2.1	18.1
		Rh <sup>3+</sup> 3d <sub>3/2</sub>	313.6	2.4	11.0
		Rh <sup>(0)</sup> 3d <sub>5/2</sub>	307.1	2.1	4.1
		Rh <sup>(0)</sup> 3d <sub>3/2</sub>	311.6	2.0	2.9
1.5 wt% Rh/ZrSnO <sub>4</sub>	0.027	Rh <sup>3+</sup> 3d <sub>5/2</sub>	309.0	2.0	18.0
		Rh <sup>3+</sup> 3d <sub>3/2</sub>	313.6	2.3	11.6
		Rh <sup>(0)</sup> 3d <sub>5/2</sub>	307.2	2.1	4.2
		Rh <sup>(0)</sup> 3d <sub>3/2</sub>	311.6	2.0	2.9
2.0 wt% Rh/ZrSnO <sub>4</sub>	0.018	Rh <sup>3+</sup> 3d <sub>5/2</sub>	309.0	2.0	18.3
		Rh <sup>3+</sup> 3d <sub>3/2</sub>	313.6	2.3	11.3
		Rh <sup>(0)</sup> 3d <sub>5/2</sub>	307.1	1.8	4.2
		Rh <sup>(0)</sup> 3d <sub>3/2</sub>	311.6	1.9	2.9

**Table 1-3.** Ratio of Rh<sup>3+</sup>/(Rh<sup>3+</sup>+Rh<sup>(0)</sup>) for  $x$  wt% Rh/ZrSnO<sub>4</sub>

Catalyst	Rh <sup>3+</sup> /(Rh <sup>3+</sup> +Rh <sup>(0)</sup> )
0.5 wt% Rh/ZrSnO <sub>4</sub>	81.5 %
1.0 wt% Rh/ZrSnO <sub>4</sub>	81.5 %
1.5 wt% Rh/ZrSnO <sub>4</sub>	81.1 %
2.0 wt% Rh/ZrSnO <sub>4</sub>	81.3 %

**Figure 1-4** displays the temperature dependence of the N<sub>2</sub>O conversion for the  $x$  wt% Rh/ZrSnO<sub>4</sub> catalysts. While the ZrSnO<sub>4</sub> solid showed the low activity, the loading of the Rh activator drastically increased the catalytic activity. For the catalysts with  $x \leq 1.0$ , with increasing the Rh loading amount ( $x$ ), the catalytic activity improved up to  $x = 1.0$ . This phenomenon can be explained by the increase in the Rh loading amount. In contrast, the decrease in the catalytic activity was observed for the catalysts with  $x > 1.0$ , likely due to the decrease in the Rh dispersion. In order to determine the effect of the Rh dispersion, the turnover frequency (TOF) values per mole of Rh at 400 °C were estimated, and the results are listed in **Table 1-4**. The TOF values decreased monotonously with increasing  $x$ . This trend is similar to that of the Rh dispersions (**Table 1-1**). The increase in the activity at  $x \leq 1.0$  was only affected by the amount of Rh loading, whereas the catalytic activity at  $x > 1.0$  decreased by the decrease in the number of active sites due to the Rh aggregation. Among the prepared catalysts, the highest activity was obtained for the 1.0 wt% Rh/ZrSnO<sub>4</sub> catalyst, and the complete N<sub>2</sub>O decomposition was achieved at 450 °C.

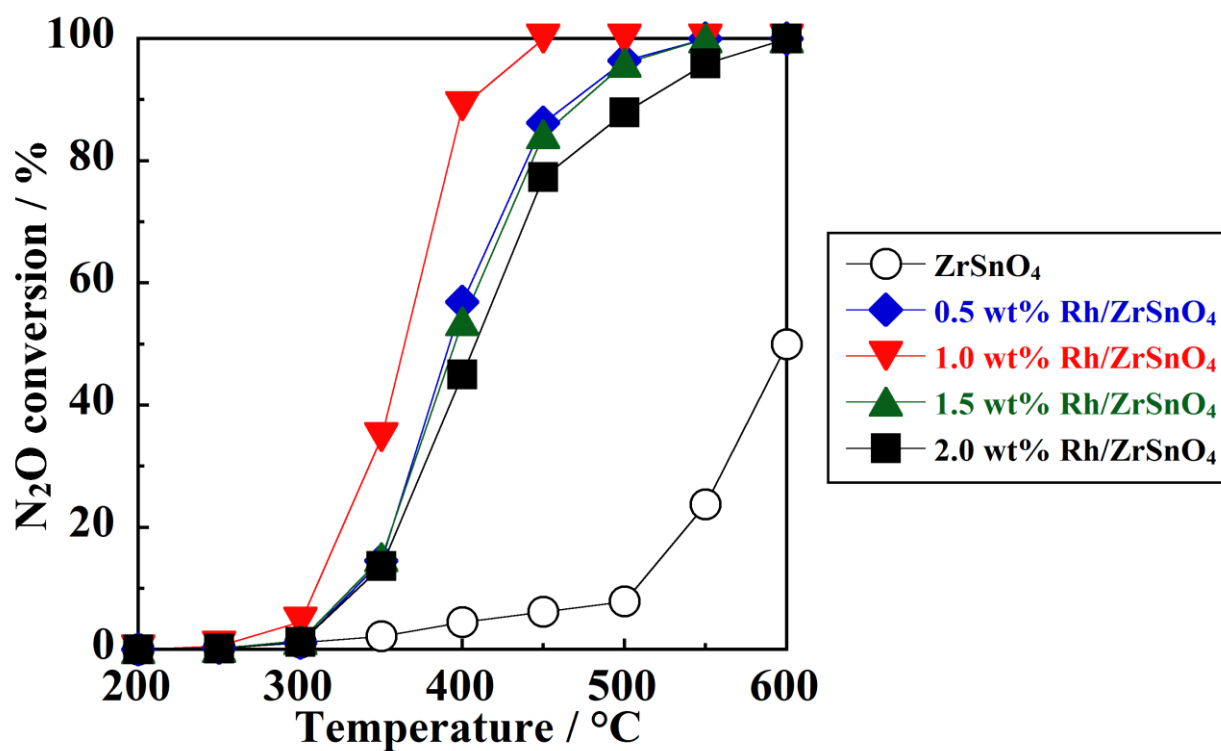
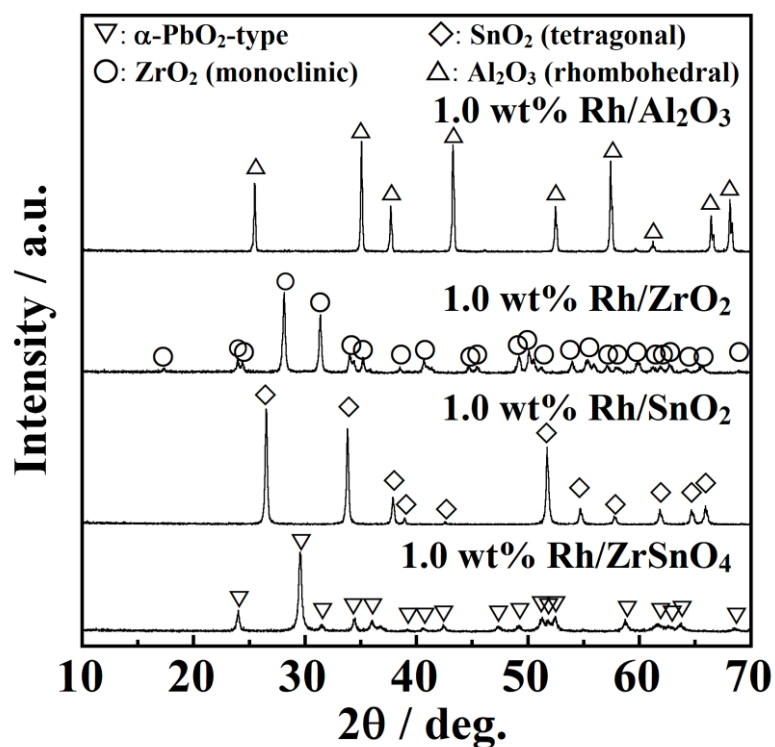


Figure 1-4. Temperature dependence of the N<sub>2</sub>O conversion for the *x* wt% Rh/ZrSnO<sub>4</sub> catalysts.

Table 1-4. TOF values per mole of Rh for the *x* wt% Rh/ZrSnO<sub>4</sub> catalysts at 400 °C

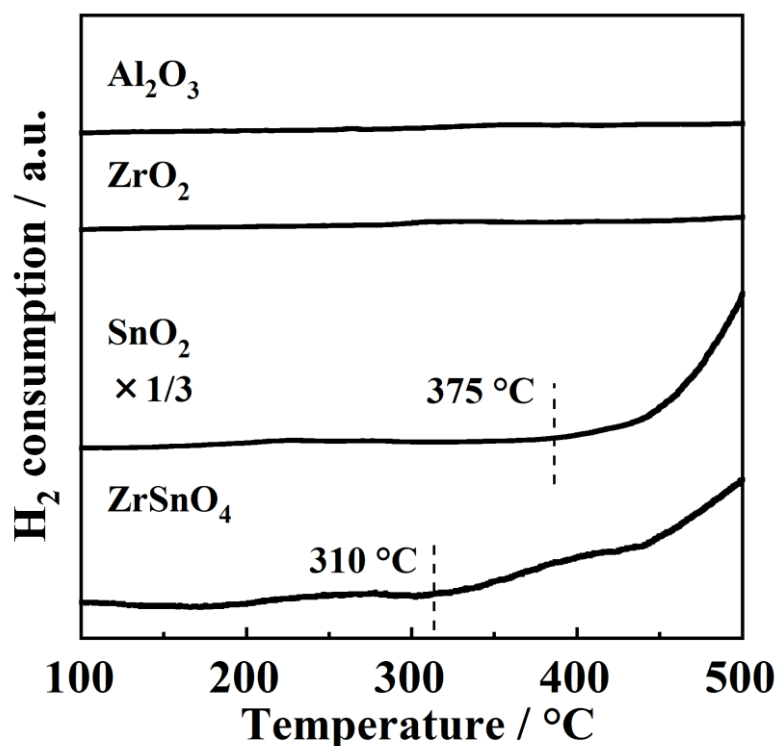
Catalyst	TOF (10 <sup>-3</sup> s <sup>-1</sup> )
0.5 wt% Rh/ZrSnO <sub>4</sub>	12.0
1.0 wt% Rh/ZrSnO <sub>4</sub>	9.4
1.5 wt% Rh/ZrSnO <sub>4</sub>	3.8
2.0 wt% Rh/ZrSnO <sub>4</sub>	2.4

In order to investigate the effect of  $\text{ZrSnO}_4$  on the catalytic activity, 1.0 wt%  $\text{Rh}/\text{SnO}_2$ , 1.0 wt%  $\text{Rh}/\text{ZrO}_2$  and 1.0 wt%  $\text{Rh}/\text{Al}_2\text{O}_3$  catalysts were prepared as the comparison. **Figure 1-5** shows the XRD patterns of 1.0 wt%  $\text{Rh}/\text{SnO}_2$ , 1.0 wt%  $\text{Rh}/\text{ZrO}_2$ , and 1.0 wt%  $\text{Rh}/\text{Al}_2\text{O}_3$ , with the data of 1.0 wt%  $\text{Rh}/\text{ZrSnO}_4$ . For each sample, no peaks related to impurities and the Rh species were detected. The specific surface areas of 1.0 wt%  $\text{Rh}/\text{SnO}_2$ , 1.0 wt%  $\text{Rh}/\text{ZrO}_2$ , and 1.0 wt%  $\text{Rh}/\text{Al}_2\text{O}_3$  were measured to be  $6.9 \text{ m}^2 \cdot \text{g}^{-1}$ ,  $13.0 \text{ m}^2 \cdot \text{g}^{-1}$ , and  $5.8 \text{ m}^2 \cdot \text{g}^{-1}$ , respectively.



**Figure 1-5.** XRD patterns of the 1.0 wt%  $\text{Rh}/\text{ZrSnO}_4$ , 1.0 wt%  $\text{Rh}/\text{SnO}_2$ , 1.0 wt%  $\text{Rh}/\text{ZrO}_2$ , and 1.0 wt%  $\text{Rh}/\text{Al}_2\text{O}_3$  catalysts.

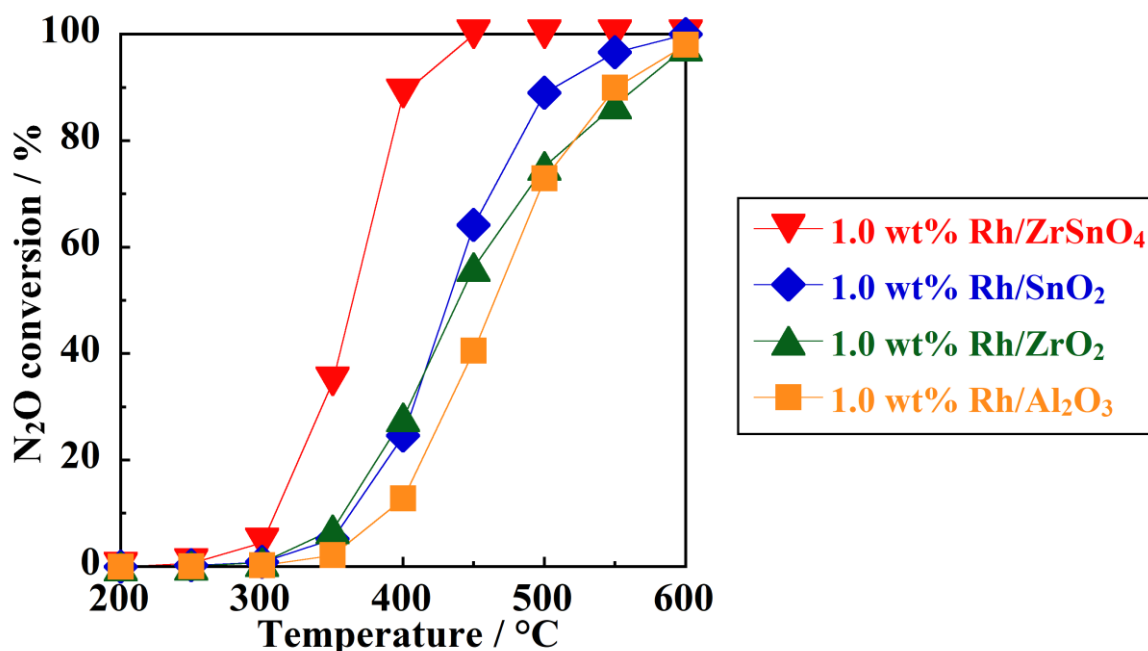
**Figure 1-6** presents the TPR profiles of the  $\text{ZrSnO}_4$ ,  $\text{SnO}_2$ ,  $\text{ZrO}_2$ , and  $\text{Al}_2\text{O}_3$  solids. For  $\text{ZrO}_2$ , and  $\text{Al}_2\text{O}_3$  solids, there were no obvious reduction peaks below 500 °C. In the cases of  $\text{ZrSnO}_4$  and  $\text{SnO}_2$ , the shoulder peak was observed at the temperatures higher than 300 °C, likely because of the reduction of  $\text{Sn}^{4+}$ . The onset reduction temperature for  $\text{ZrSnO}_4$  was 310 °C, which is lower than that for  $\text{SnO}_2$  (375 °C). Therefore,  $\text{ZrSnO}_4$  can release oxygen species from inside the lattice, and its ability is higher than those for  $\text{SnO}_2$ ,  $\text{ZrO}_2$ , and  $\text{Al}_2\text{O}_3$ .



**Figure 1-6.** TPR profiles of the  $\text{ZrSnO}_4$ ,  $\text{SnO}_2$ ,  $\text{ZrO}_2$ , and  $\text{Al}_2\text{O}_3$  solids.



**Figure 1-7** shows the temperature dependence of the N<sub>2</sub>O conversion for the 1.0 wt% Rh/ZrSnO<sub>4</sub>, 1.0 wt% Rh/SnO<sub>2</sub>, 1.0 wt% Rh/ZrO<sub>2</sub>, and 1.0 wt% Rh/Al<sub>2</sub>O<sub>3</sub> catalysts. Regardless of the same Rh loading amount and similar low specific surface areas (1.0 wt% Rh/ZrSnO<sub>4</sub>: 13.6 m<sup>2</sup>·g<sup>-1</sup>, 1.0 wt% Rh/SnO<sub>2</sub>: 6.9 m<sup>2</sup>·g<sup>-1</sup>, 1.0 wt% Rh/ZrO<sub>2</sub>: 13.0 m<sup>2</sup>·g<sup>-1</sup>, 1.0 wt% Rh/Al<sub>2</sub>O<sub>3</sub>: 5.8 m<sup>2</sup>·g<sup>-1</sup>), the catalytic activity of 1.0 wt% Rh/ZrSnO<sub>4</sub> was significantly higher than those of 1.0 wt% Rh/SnO<sub>2</sub>, 1.0 wt% Rh/ZrO<sub>2</sub>, and 1.0 wt% Rh/Al<sub>2</sub>O<sub>3</sub>. The highest oxygen release ability of ZrSnO<sub>4</sub> might contribute to the highest catalytic activity. Here, the activity of 1.0 wt% Rh/SnO<sub>2</sub> was higher than those of 1.0 wt% Rh/ZrO<sub>2</sub> and 1.0 wt% Rh/Al<sub>2</sub>O<sub>3</sub>, because of the oxygen release from the SnO<sub>2</sub> lattice. Therefore, the ZrSnO<sub>4</sub> solid effectively promoted the N<sub>2</sub>O decomposition over Rh, because the oxygen supply from inside the ZrSnO<sub>4</sub> lattice toward the Rh activator might remove adsorbed oxygen species which were generated on Rh during the N<sub>2</sub>O decomposition process.



**Figure 1-7.** Temperature dependence of the N<sub>2</sub>O conversion for the 1.0 wt% Rh/ZrSnO<sub>4</sub>, 1.0 wt% Rh/SnO<sub>2</sub>, 1.0 wt% Rh/ZrO<sub>2</sub>, and 1.0 wt% Rh/Al<sub>2</sub>O<sub>3</sub> catalysts.

## 1.4. Conclusion

In this chapter, in order to develop novel catalysts for  $\text{N}_2\text{O}$  decomposition, Rh supported on  $\text{ZrSnO}_4$  were synthesized. Among the prepared catalysts, the 1.0 wt% Rh/ $\text{ZrSnO}_4$  catalyst exhibited the highest activity, and  $\text{N}_2\text{O}$  was completely decomposed at 450 °C. In addition, Rh/ $\text{ZrSnO}_4$  exhibited the higher catalytic activity than those of Rh/ $\text{SnO}_2$ , Rh/ $\text{ZrO}_2$ , and Rh/ $\text{Al}_2\text{O}_3$ , suggesting that the oxygen supply from the  $\text{ZrSnO}_4$  lattice could facilitate the removal of adsorbed surface oxygen, generated on the Rh activator during the  $\text{N}_2\text{O}$  decomposition process.

## ***Chapter 2***

### **Noble-metal-free Catalysts Based on $\text{ZrSnO}_4$ for Direct Decomposition of $\text{N}_2\text{O}$**

#### **2.1. Introduction**

Chapter 1 demonstrated that  $\text{ZrSnO}_4$  could promote  $\text{N}_2\text{O}$  decomposition over the Rh activator, and 1.0 wt% Rh/ $\text{ZrSnO}_4$  completely decomposed  $\text{N}_2\text{O}$  at 450 °C. However, owing to the high cost and scarce resources of noble metals, noble-metal-free catalysts are desired for the widespread industrial applications.

This chapter aims the development of novel noble-metal-free catalysts with high activity. To this end,  $\text{Co}_3\text{O}_4$  was selected as an activator instead of Rh, because spinel-type  $\text{Co}_3\text{O}_4$  based catalysts are known to be effective for  $\text{N}_2\text{O}$  decomposition, described in the General Introduction section. Furthermore, in order to increase the oxygen supply ability of  $\text{ZrSnO}_4$ ,  $\text{Fe}^{2+/3+}$  ions were introduced into  $\text{ZrSnO}_4$ , where the valence change of  $\text{Fe}^{2+/3+}$  may improve the redox properties and the replacement of the  $\text{Sn}^{4+}$  site for lower-valent  $\text{Fe}^{2+/3+}$  can generate oxygen vacancies for smooth oxide ion migration. Therefore, noble-metal-free catalysts of  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  were synthesized, and the activities for  $\text{N}_2\text{O}$  direct decomposition were investigated.

#### **2.2. Experimental Procedure**

The  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  powders were prepared by a co-precipitation method by using starting

materials of zirconium (IV) oxynitrate dihydrate ( $\text{ZrO}(\text{NO}_3)_2 \cdot 2\text{H}_2\text{O}$ , Kishida Chemical,  $\geq 99.0\%$ ), tin (II) oxalate ( $\text{SnC}_2\text{O}_4$ , Sigma-Aldrich, 99%), and  $0.5 \text{ mol} \cdot \text{L}^{-1} \text{Fe}(\text{NO}_3)_3$  aqueous solution (prepared by dissolving iron (III) nitrate nonahydrate ( $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ , FUJIFILM Wako Pure Chemical, 99.9%) in deionized water) according to a Zr:Sn:Fe molar ratio of  $1.3:(1-x):x$ . The subsequent experimental methods are the same as described in Chapter 1.

The  $\text{Co}_3\text{O}_4$  loading was conducted by impregnating the  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  powders with cobalt (II) nitrate hexahydrate ( $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , Kishida Chemical,  $\geq 98.0\%$ ) in 40 mL of deionized water. The 1.1 molar equivalents to Co ions of citric acid (FUJIFILM Wako Pure Chemical, 98.0%) and 5.0 molar equivalents to Co ions of 1,2-propanediol (Kishida Chemical,  $\geq 99.0\%$ ) were added as dispersants. After the mixture was stirred at  $90^\circ\text{C}$  for 4 h, the solvent was evaporated at  $150^\circ\text{C}$ . Subsequently, the resulting solid was pulverized, and then calcinated at  $600^\circ\text{C}$  for 1 h in the air. The  $\text{Co}_3\text{O}_4$  loading amount was adjusted to be  $y \text{ wt}\%$  ( $y = 11, 18, 19, 20, \text{ and } 27$ ). In particular, 19 wt%  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  were denoted as  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ .

The obtained sample compositions were determined by using XRF (Supermini200, Rigaku). The BET specific surface area was determined by using a Micromeritics TriStar 3000 analyzer (Shimadzu). The crystal structures were identified through XRD (SmartLab, Rigaku) with  $\text{Cu K}\alpha$  radiation, where the lattice volumes were calculated through refining the XRD peak angles using  $\alpha$ -alumina as the internal standard. TPR measurements were conducted as described in Chapter 1. The catalytic activity for  $\text{N}_2\text{O}$  direct decomposition was also carried out according to Chapter 1.

### 2.3. Results and Discussion

The XRD patterns of the  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  ( $x = 0.03, 0.05, 0.07, 0.10$ ) solids are presented in **Figure 2-1**. All the samples showed a single-phase of the  $\alpha\text{-PbO}_2$ -type structure, where no crystalline phases were detected. The lattice volumes of  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  were estimated using the XRD patterns, and the compositional dependence of the lattice volume is depicted in **Figure 2-2**. With increasing the  $\text{Fe}^{2+/3+}$  content ( $x$ ), the lattice volume decreased monotonously until  $x = 0.07$ , indicating that the  $\text{Sn}^{4+}$  ( $\text{Sn}^{4+}$ : 0.083 nm, coordination number [CN] = 6) site in  $\text{ZrSnO}_4$  was partially replaced for the smaller  $\text{Fe}^{2+/3+}$  ions ( $\text{Fe}^{2+}$ : 0.075 nm,  $\text{Fe}^{3+}$ : 0.069 nm, CN = 6) [37]. Here, since the valence states of  $\text{Fe}^{2+/3+}$  are lower than that of  $\text{Sn}^{4+}$ , oxygen vacancies should be formed in the lattice by the charge compensation mechanism. For the sample with  $x = 0.10$ , no significant lattice volume change was observed compared to the sample with  $x = 0.07$ . Since the  $x = 0.10$  sample showed only the peaks corresponding to the  $\alpha\text{-PbO}_2$ -type phase from the XRD pattern, an excess Fe component might exist as an amorphous phase. Therefore, the solid solubility limit composition is considered to be  $\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  ( $x = 0.07$ ).

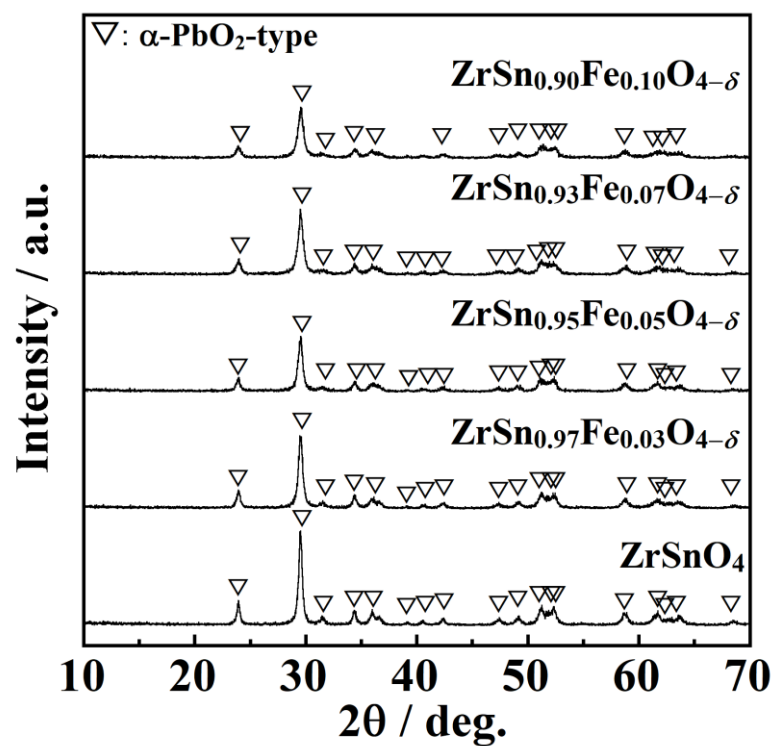


Figure 2-1. XRD patterns of the  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  solids.

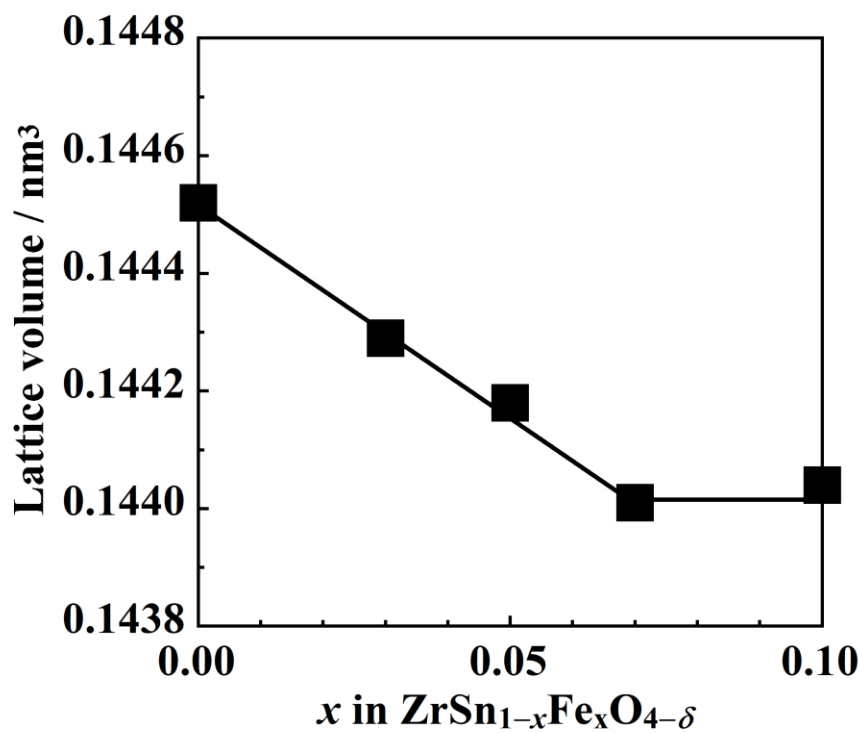


Figure 2-2. Compositional dependence of the lattice volume for  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ .

In order to investigate the effect of the  $\text{Fe}^{2+/3+}$  introduction on the redox properties of  $\text{ZrSnO}_4$ , TPR measurement was performed for  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ , and the TPR profiles are shown in **Figure 2-3**. For  $\text{ZrSnO}_4$ , there were no obvious reduction peaks below 300 °C, whereas the shoulder peak was observed at the temperatures higher than 300 °C. In the cases of  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ , all the samples have one broad reduction peak (ca. 270 °C) below 300 °C, which were identified as the reduction of  $\text{Fe}^{3+}$  [38]. In the range of  $x \leq 0.07$ , the reduction peaks shifted toward lower temperatures with increasing the Fe content ( $x$ ), and the lowest reduction temperature was obtained for the sample with  $x = 0.07$  (259 °C). This acceleration of the oxygen release can be explained by the introduction of  $\text{Fe}^{2+/3+}$  into  $\text{ZrSnO}_4$ , which improved the redox properties owing to the valence change of  $\text{Fe}^{2+/3+}$  and generated oxygen vacancies for smooth oxide ion conduction. For  $x = 0.10$ , the reduction temperature shifted toward higher temperatures, due to amorphous phase of excess Fe component on the surface. Therefore, it is considered that the  $x = 0.07$  sample exhibited the highest oxygen supply ability.

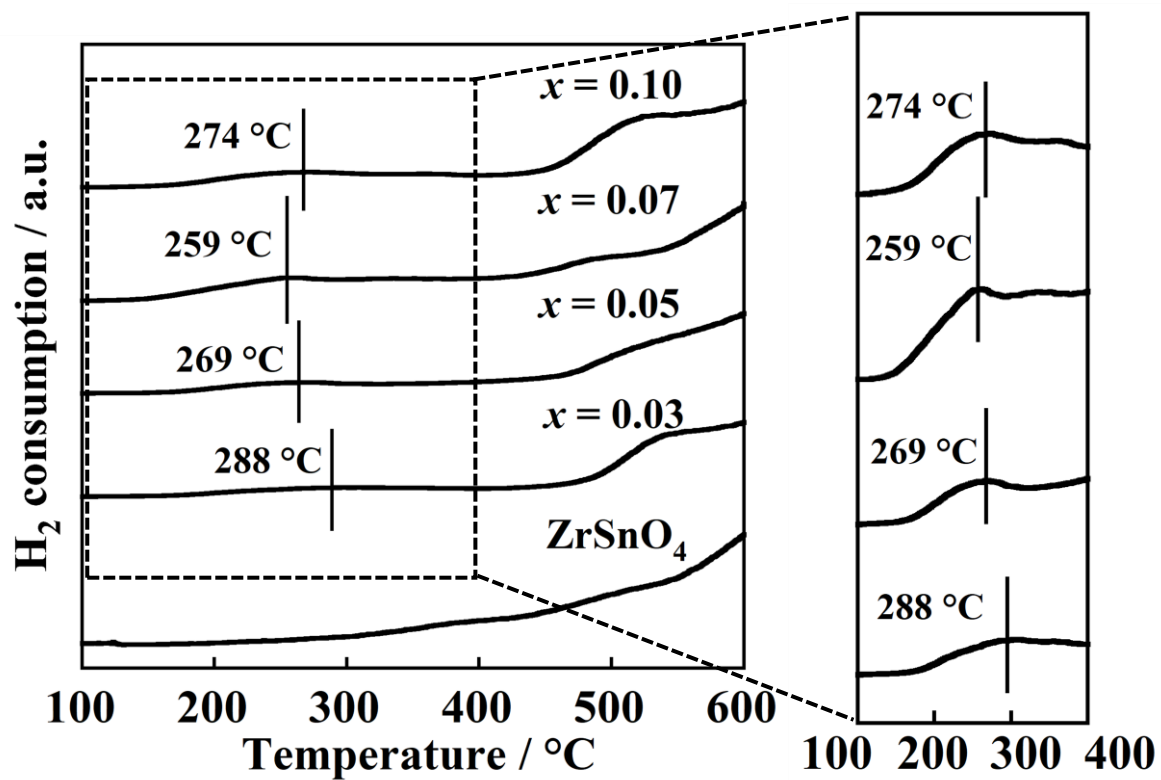
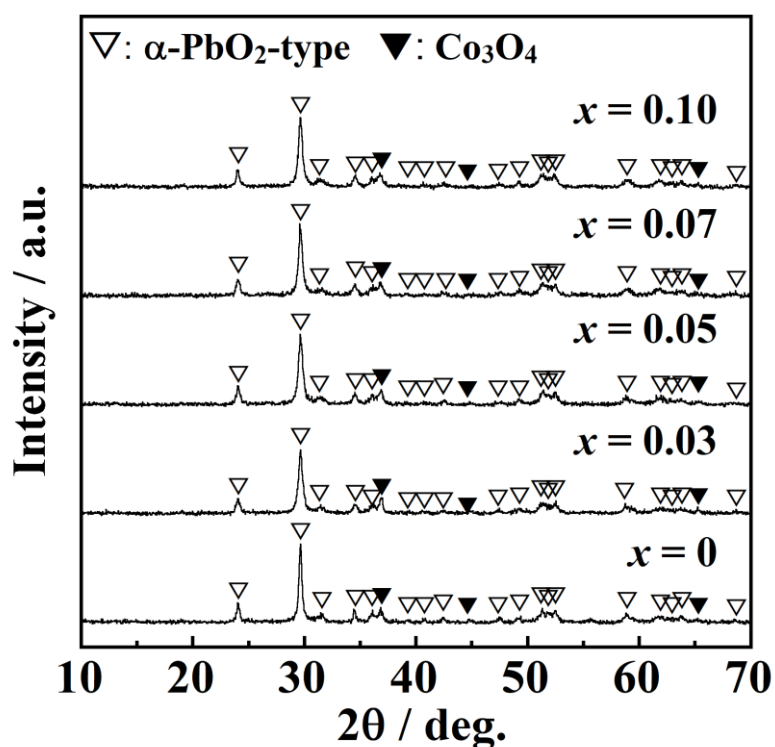


Figure 2-3. TPR profiles of  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ .



**Figure 2-4** depicts the XRD patterns of  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ . For all the samples, the peaks were assigned to be the  $\alpha\text{-PbO}_2$ -type phase and the spinel-type  $\text{Co}_3\text{O}_4$  phase, suggesting that  $\text{Co}_3\text{O}_4$  was supported onto  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ . **Table 2-1** shows the crystallite sizes of  $\text{Co}_3\text{O}_4$  and  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  estimated from the XRD results. For all the samples, the crystallite sizes of  $\text{Co}_3\text{O}_4$  (ca. 7 nm) and  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  (ca. 26 nm) were almost the same, irrespective of the  $\text{Fe}^{2+/3+}$  content ( $x$ ).

**Table 2-2** shows the measured compositions and the specific surface areas of the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts. The measured compositions were in good agreement with their feed values, within the experimental error. In addition, the specific surface areas of all the samples were similar (ca.  $31 \text{ m}^2 \cdot \text{g}^{-1}$ ).



**Figure 2-4.** XRD patterns of the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts.

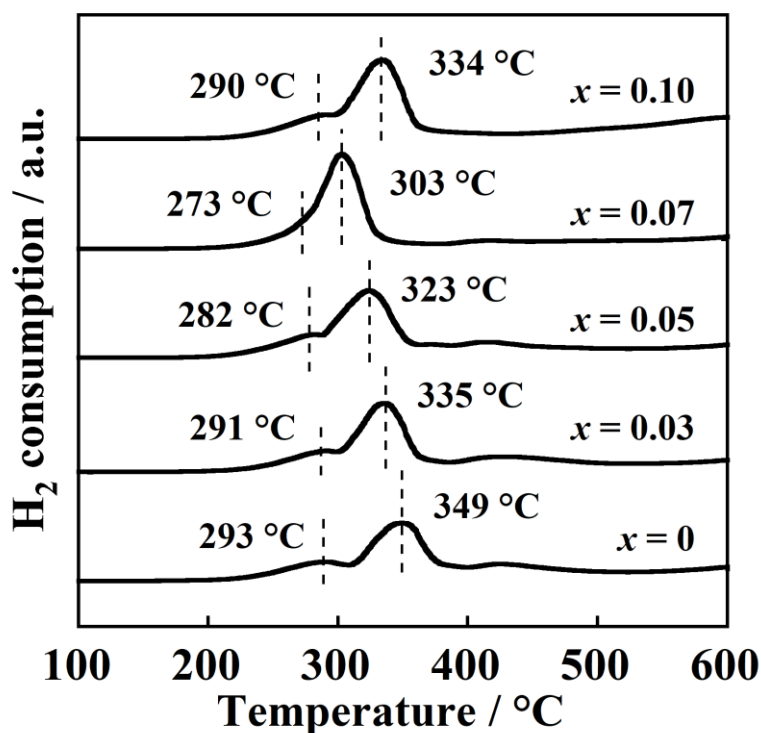
**Table 2-1.** Crystallite sizes of  $\text{Co}_3\text{O}_4$  and  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  for the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts

Catalyst	$\text{Co}_3\text{O}_4$ / nm	$\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$ / nm
$x = 0$	7	30
$x = 0.03$	7	23
$x = 0.05$	6	24
$x = 0.07$	7	27
$x = 0.10$	8	24

**Table 2-2.** Measured composition and specific surface area of the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts

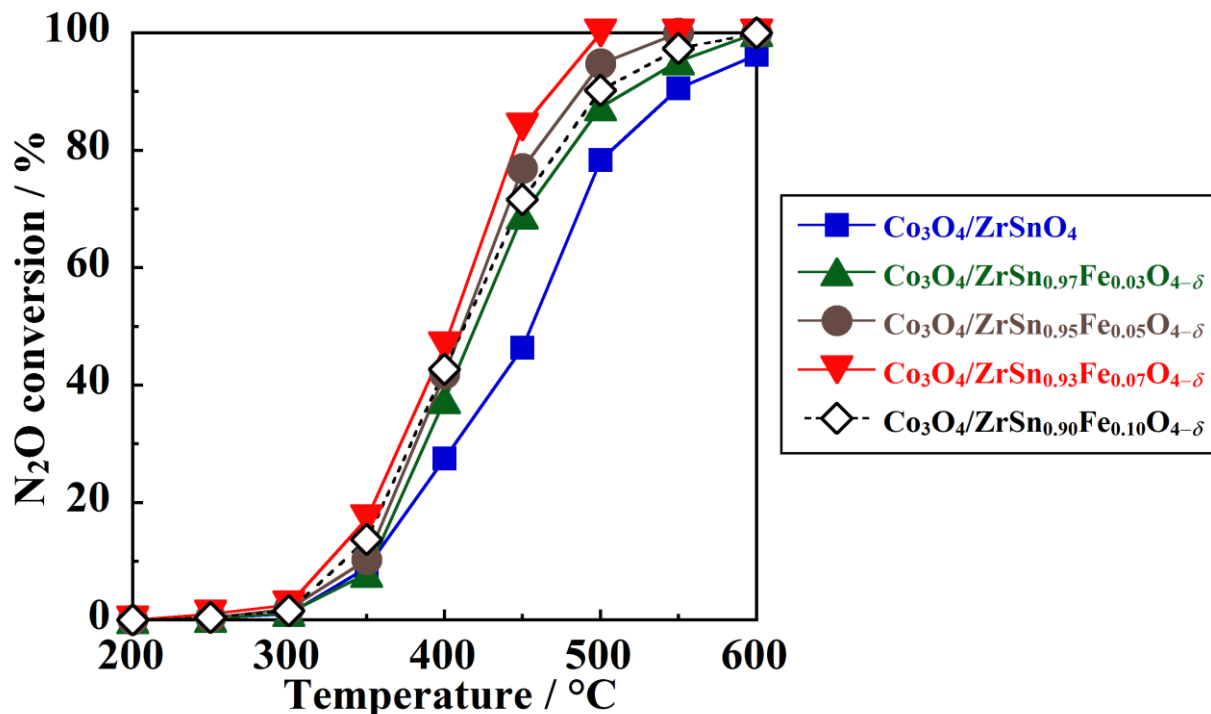
Feed composition	Measured composition	Specific surface area / $\text{m}^2 \cdot \text{g}^{-1}$
19wt% $\text{Co}_3\text{O}_4/\text{ZrSnO}_4$	19.3wt% $\text{Co}_3\text{O}_4/\text{Zr}_{1.05}\text{Sn}_{0.95}\text{O}_4$	31.6
19wt% $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.97}\text{Fe}_{0.03}\text{O}_{4-\delta}$	19.7wt% $\text{Co}_3\text{O}_4/\text{Zr}_{1.04}\text{Sn}_{0.93}\text{Fe}_{0.03}\text{O}_{4-\delta}$	31.4
19wt% $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.95}\text{Fe}_{0.05}\text{O}_{4-\delta}$	19.5wt% $\text{Co}_3\text{O}_4/\text{Zr}_{1.06}\text{Sn}_{0.89}\text{Fe}_{0.05}\text{O}_{4-\delta}$	30.4
19wt% $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$	19.2wt% $\text{Co}_3\text{O}_4/\text{Zr}_{1.07}\text{Sn}_{0.86}\text{Fe}_{0.07}\text{O}_{4-\delta}$	31.4
19wt% $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.90}\text{Fe}_{0.10}\text{O}_{4-\delta}$	19.2wt% $\text{Co}_3\text{O}_4/\text{Zr}_{1.07}\text{Sn}_{0.84}\text{Fe}_{0.10}\text{O}_{4-\delta}$	32.5

**Figure 2-5** presents the TPR profiles of the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts. The  $\text{Co}_3\text{O}_4/\text{ZrSnO}_4$  catalyst exhibited two peaks at 293 °C and 349 °C, assigned as the reduction of surface  $\text{Co}^{3+}$  and bulk  $\text{Co}^{3+}$ , respectively [39]. These reduction peaks shifted toward lower temperatures with increasing the  $\text{Fe}^{2+/3+}$  content for  $x \leq 0.07$ . This easier reduction of  $\text{Co}_3\text{O}_4$  may be attributed to the improved oxygen supply ability of the  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  solid. For the  $x = 0.10$  catalyst, each reduction peak shifted toward higher temperatures compared to that for the  $x = 0.07$  catalyst, which was related to the Fe based amorphous phase on the catalyst surface. From these results, the  $x = 0.07$  catalyst showed the highest reducibility; that is, the oxygen release from  $\text{Co}_3\text{O}_4$  was facilitated by the oxygen supply from the  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  lattice.



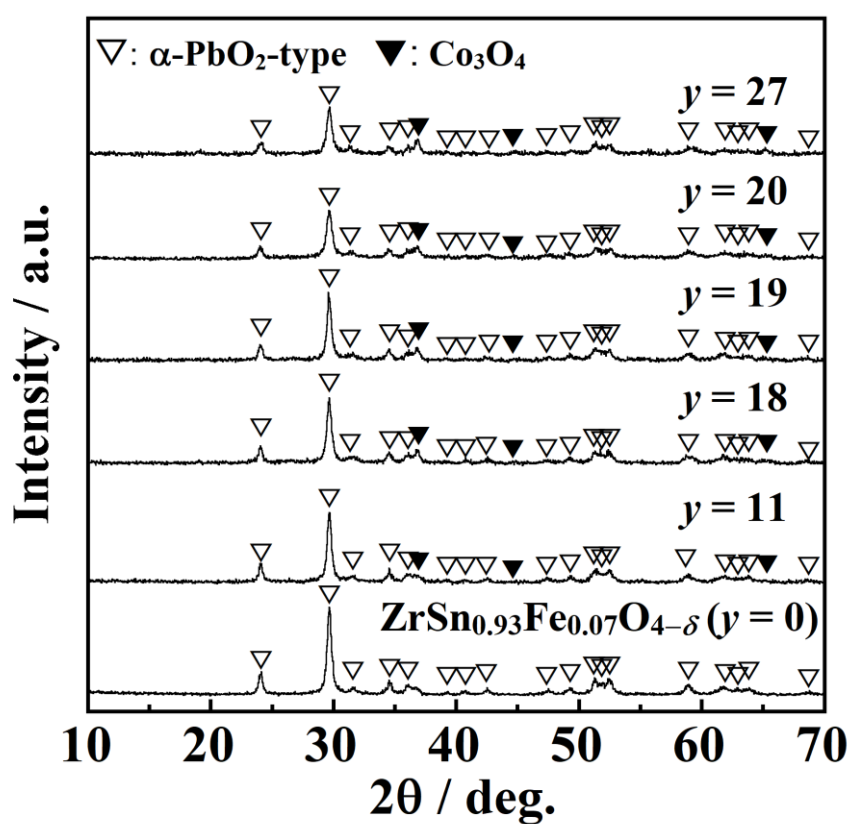
**Figure 2-5.** TPR profiles of the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts.

The temperature dependence of the  $\text{N}_2\text{O}$  conversion for  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  are depicted in **Figure 2-6**. With the increase in  $x$ , the catalytic activity of the catalysts with  $x \leq 0.07$  obviously enhanced, because the effective oxygen supply from inside the  $\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  lattice toward  $\text{Co}_3\text{O}_4$  facilitated the removal of adsorbed oxygen. On the other hand, the catalyst with  $x = 0.10$  showed lower activity than that for  $x = 0.07$ , because of its decreased oxygen supply ability. Therefore, the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  catalyst exhibited the highest activity, where  $\text{N}_2\text{O}$  was completely decomposed at 500 °C. Although the  $\text{Rh}/\text{ZrSnO}_4$  catalyst could decompose  $\text{N}_2\text{O}$  completely at 450 °C (as described in Chapter 1), it relies on the use of the noble metal Rh. The developed  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  catalyst exhibited a high  $\text{N}_2\text{O}$  decomposition efficiency of 84% at the same temperature of 450 °C without using the noble metal.



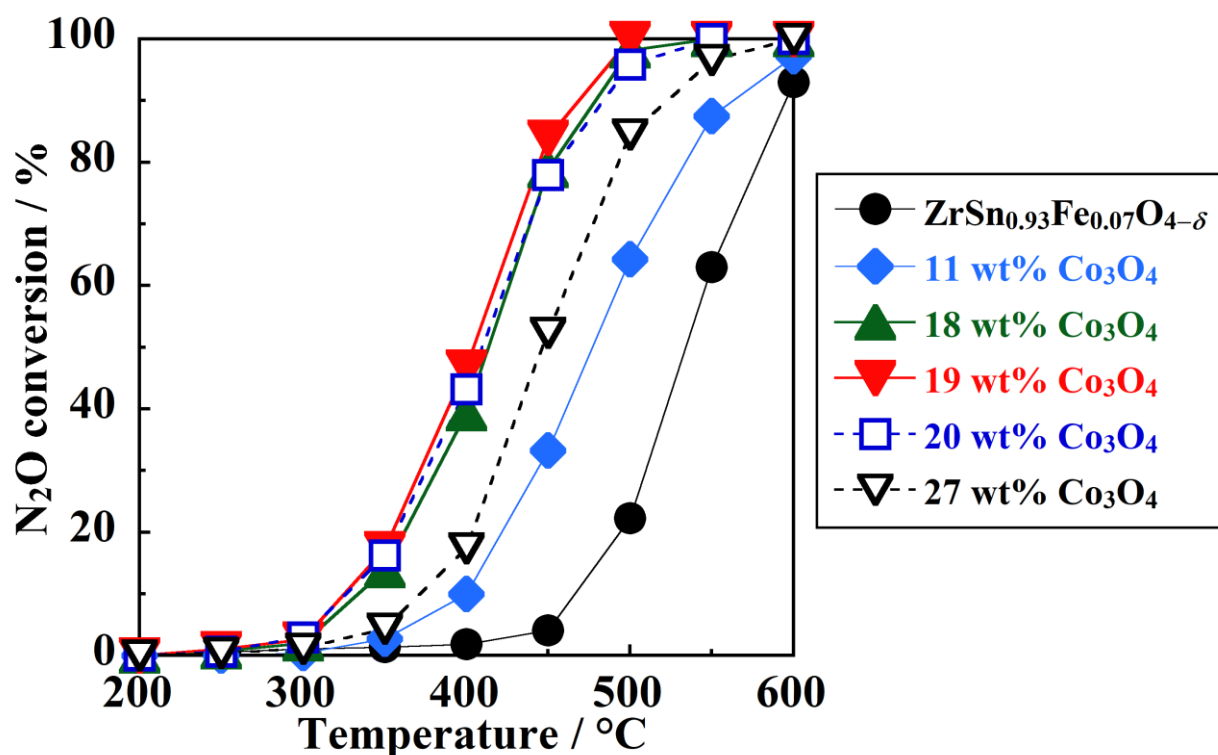
**Figure 2-6.** Temperature dependence of the  $\text{N}_2\text{O}$  conversion for the  $\text{Co}_3\text{O}_4/\text{ZrSn}_{1-x}\text{Fe}_x\text{O}_{4-\delta}$  catalysts.

The effect of the  $\text{Co}_3\text{O}_4$  loading amount on the catalytic activity was investigated. The XRD patterns of the  $y$  wt%  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  ( $y = 11, 18, 19, 20, 27$ ) catalysts are shown in **Figure2-7**. For each sample, peaks identified as the  $\alpha\text{-PbO}_2$ -type structure and the  $\text{Co}_3\text{O}_4$  phase were observed without any peaks caused by impurities.



**Figure 2-7.** XRD patterns of the  $y$  wt%  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  catalysts.

**Figure 2-8** shows the temperature dependence of the  $\text{N}_2\text{O}$  conversion for the  $y$  wt%  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  catalysts. With increasing the  $\text{Co}_3\text{O}_4$  loading amount ( $y$ ), the catalytic activity improved up to  $y = 19$ . This phenomenon can be explained by the increase in the amount of the  $\text{Co}_3\text{O}_4$  activator. For the catalysts with  $y > 19$ , the catalytic activity decreased probably due to the aggregation of  $\text{Co}_3\text{O}_4$ . Therefore, the loading amount of the  $\text{Co}_3\text{O}_4$  activator was optimized to be 19 wt%.



**Figure 2-8.** Temperature dependence of the  $\text{N}_2\text{O}$  conversion for the  $y$  wt%  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  catalysts.

## 2.4. Conclusion

In this chapter, in order to develop noble-metal-free catalysts with high catalytic activity for N<sub>2</sub>O direct decomposition, the Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>1-x</sub>Fe<sub>x</sub>O<sub>4-δ</sub> catalysts were prepared. The introduction of Fe<sup>2+/3+</sup> ions into the ZrSnO<sub>4</sub> lattice enhanced the N<sub>2</sub>O decomposition activity, because of the redox properties of the Fe<sup>2+/3+</sup> couple and the generation of oxygen vacancies for the smooth oxide ion conduction. Among the prepared catalysts, the 19 wt% Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> catalyst exhibited the highest activity and achieved complete decomposition of N<sub>2</sub>O at 500 °C without using noble metals.

## Chapter 3

### ZrSnO<sub>4</sub> Based Catalysts with Durability against Oxygen Gas and Water Vapor

#### 3.1. Introduction

In Chapter 2, 19 wt% Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> completely decomposed N<sub>2</sub>O at 500 °C, while it does not contain noble metals. However, the spinel-type Co<sub>3</sub>O<sub>4</sub> catalyst is reported to have low durability against coexisting gases of oxygen gas and water vapor [40]. Here, practically applied catalysts are often exposed to coexisting gases; e.g. 4 vol% O<sub>2</sub> and 2–3 vol% H<sub>2</sub>O vapor during adipic acid and nitric acid production [41]. Therefore, the development of catalysts with high durability against coexisting gases of oxygen gas and water vapor is desired.

This chapter aims the development of novel catalysts with high durability against coexisting gases of O<sub>2</sub> and H<sub>2</sub>O vapor. To this end, the doping of Co<sup>2+/3+</sup> was carried out into the ZrSnO<sub>4</sub> lattice, instead of the loading of Co<sup>2+/3+</sup> as Co<sub>3</sub>O<sub>4</sub> with low durability. The doping of Co<sup>2+/3+</sup> was expected to improve the redox properties due to two types of valence states of Co<sup>2+</sup> and Co<sup>3+</sup> and to increase oxygen vacancies through the replacement of the Sn<sup>4+</sup> site for low-valent Co<sup>2+/3+</sup>. In addition, since the ionic-size of Sn<sup>4+</sup> was closer to Co<sup>2+</sup> than to Fe<sup>2+</sup> (Sn<sup>4+</sup>: 0.083 nm, Co<sup>2+</sup>: 0.079 nm, Fe<sup>2+</sup>: 0.075 nm, CN = 6) [37], the high amount of dopant is expected to be introduced into the Sn<sup>4+</sup> site in the ZrSnO<sub>4</sub> lattice, compared to the case of the Fe<sup>2+/3+</sup> doped ZrSnO<sub>4</sub> series. Therefore, ZrSn<sub>1-x</sub>Co<sub>x</sub>O<sub>4-δ</sub> catalysts were prepared, and their N<sub>2</sub>O direct decomposition activities and durability against coexisting gases of O<sub>2</sub> and H<sub>2</sub>O vapor were investigated.



### 3.2. Experimental Procedure

The  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  powders were prepared by a co-precipitation method by using starting materials of zirconium (IV) oxynitrate dihydrate ( $\text{ZrO}(\text{NO}_3)_2 \cdot 2\text{H}_2\text{O}$ , Kishida Chemical,  $\geq 99.0\%$ ) powder, tin (II) oxalate ( $\text{SnC}_2\text{O}_4$ , Sigma-Aldrich, 99%) powder, and cobalt (II) nitrate hexahydrate ( $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , Kishida Chemical,  $\geq 98.0\%$ ) powder, where a feed Zr:Sn:Co molar ratio was 1.3:(1- $x$ ): $x$ . The subsequent experimental methods are the same as described in Chapter 1.

The obtained sample compositions were determined by using XRF (Supermini200, Rigaku). The BET specific surface area was determined by using a Micromeritics TriStar 3000 analyzer (Shimadzu). The crystal structures were identified through XRD (SmartLab, Rigaku) with Cu  $K\alpha$  radiation. The lattice volumes were calculated through refining the XRD peak angles using  $\alpha$ -alumina as the internal standard. TPR measurements were conducted under a reducing gas flow (5 vol%  $\text{H}_2$  - 95 vol% Ar at  $50 \text{ mL} \cdot \text{min}^{-1}$ ) at a heating rate of  $5 \text{ }^\circ\text{C} \cdot \text{min}^{-1}$  over 0.035 g of catalysts (BELCAT-B, MicrotracBEL). The oxidation states of Co species were analyzed by XPS (PHI5000 VersaProbe II, ULVAC-PHI) analysis with Al  $K\alpha$  radiation, where the effect of charging on the binding energies was refined by the C 1s peak (284.6 eV). The XPS spectrum was fitted using a Shirley background and Gaussian–Lorentzian line shapes.

The catalytic activity for  $\text{N}_2\text{O}$  direct decomposition was carried out according to Chapter 1. The effect of the coexistence of  $\text{O}_2$  and  $\text{H}_2\text{O}$  vapor on the activity was investigated by mixing  $\text{O}_2$  gas with the reactant gas (0.5 vol%  $\text{N}_2\text{O}$  - 99.5 vol% He) and bubbling the gas into deionized water. During the durability test, the total flow rate of the mixed gas was set at  $60 \text{ mL} \cdot \text{min}^{-1}$ .

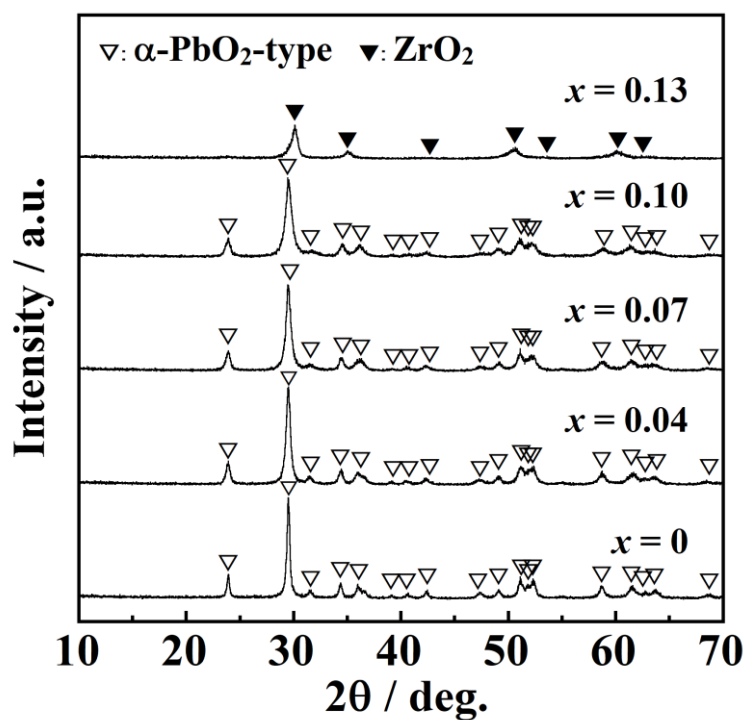
### 3.3. Results and Discussion

The measured compositions of the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  catalysts are listed in **Table 3-1**. For all the samples, the Co content in the measured composition was lower than that in the feed composition, probably due to the generation of a water-soluble complex (e.g.,  $\text{Co}(\text{NH}_3)_6^{3+}$ ) during the co-precipitation process. Nevertheless, an almost linear relationship was obtained between the Co content in the measured values and that in the feed values. Hereafter, the Co contents ( $x$ ) were expressed using measured values; that is,  $\text{Zr}_{1.07}\text{Sn}_{0.93}\text{O}_4$ ,  $\text{Zr}_{1.09}\text{Sn}_{0.87}\text{Co}_{0.04}\text{O}_{4-\delta}$ ,  $\text{Zr}_{1.15}\text{Sn}_{0.78}\text{Co}_{0.07}\text{O}_{4-\delta}$ ,  $\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$ , and  $\text{Zr}_{1.24}\text{Sn}_{0.63}\text{Co}_{0.13}\text{O}_{4-\delta}$ , were denoted as  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  ( $x = 0, 0.04, 0.07, 0.10$ , and  $0.13$ ), respectively, in Chapter 3. The BET specific surface areas of the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  catalysts are also listed in **Table 3-1**. All the samples showed similar specific surface areas of ca.  $30 \text{ m}^2 \cdot \text{g}^{-1}$ .

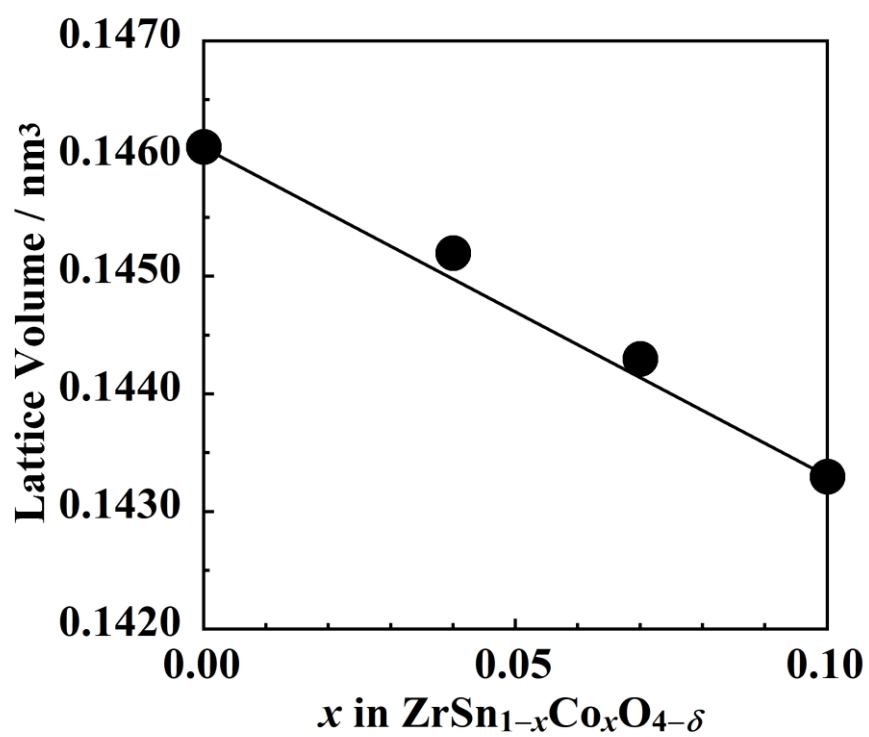
**Table 3-1.** Measured composition and specific surface area of the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  catalysts

Feed composition	Measured composition	Specific surface area / $\text{m}^2 \cdot \text{g}^{-1}$
$\text{ZrSnO}_4$	$\text{Zr}_{1.07}\text{Sn}_{0.93}\text{O}_4$	30.2
$\text{ZrSn}_{0.90}\text{Co}_{0.10}\text{O}_{4-\delta}$	$\text{Zr}_{1.09}\text{Sn}_{0.87}\text{Co}_{0.04}\text{O}_{4-\delta}$	29.6
$\text{ZrSn}_{0.80}\text{Co}_{0.20}\text{O}_{4-\delta}$	$\text{Zr}_{1.15}\text{Sn}_{0.78}\text{Co}_{0.07}\text{O}_{4-\delta}$	27.2
$\text{ZrSn}_{0.70}\text{Co}_{0.30}\text{O}_{4-\delta}$	$\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$	32.2
$\text{ZrSn}_{0.60}\text{Co}_{0.40}\text{O}_{4-\delta}$	$\text{Zr}_{1.24}\text{Sn}_{0.63}\text{Co}_{0.13}\text{O}_{4-\delta}$	29.7

The XRD patterns of the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  ( $x = 0.04, 0.07, 0.10, 0.13$ ) catalysts are shown in **Figure 3-1**. The catalysts with  $x \leq 0.10$  exhibited a single-phase  $\alpha\text{-PbO}_2$ -type structure. For the catalyst with  $x = 0.13$ , no peak corresponding to the  $\alpha\text{-PbO}_2$ -type phase was observed, and only the cubic  $\text{ZrO}_2$  phase was detected. **Figure 3-2** displays the compositional dependence of the lattice volume for the catalysts with  $x \leq 0.10$ . With increasing  $x$ , the lattice volume of the  $\alpha\text{-PbO}_2$ -type phase decreased monotonously with increasing the  $\text{Co}^{2+/3+}$  content ( $x$ ) up to  $x = 0.10$ , indicating that the  $\text{Sn}^{4+}$  site (0.083 nm, CN = 6) [37] was successfully replaced for the smaller  $\text{Co}^{2+/3+}$  ions ( $\text{Co}^{2+}$ : 0.079 nm,  $\text{Co}^{3+}$ : 0.069 nm, CN = 6) [37]. Here, as shown in **Table 3-1**, the content of  $\text{Zr}^{4+}$  having large ionic-size (0.086 nm, CN = 6) [37] was also increased with increasing the  $\text{Co}^{2+/3+}$  content, which supports that the lattice shrinkage was mainly affected by the introduction of the small ionic-size of  $\text{Co}^{2+/3+}$ .

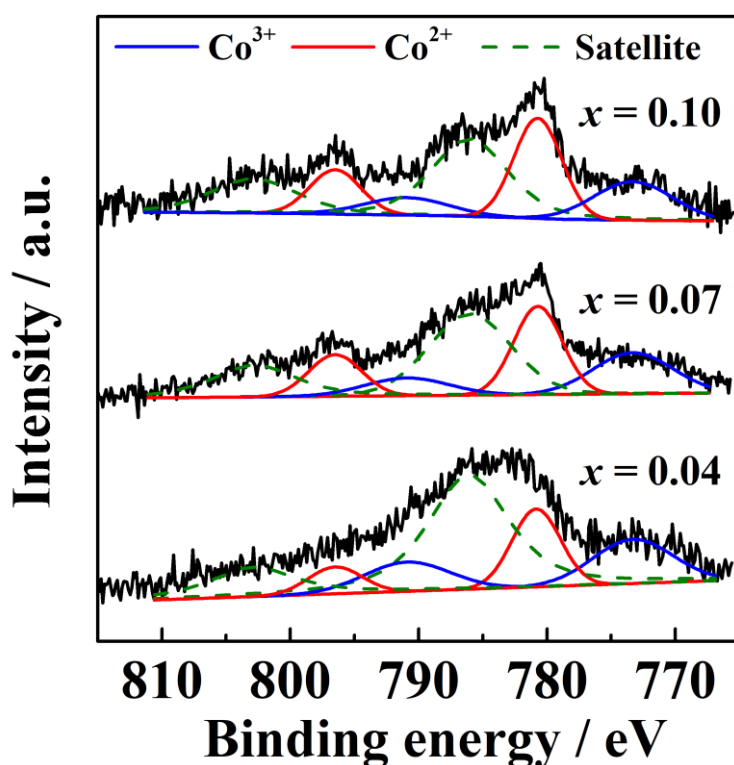


**Figure 3-1.** XRD patterns of the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  catalysts.



**Figure 3-2.** Compositional dependence of the lattice volume for  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$ .

**Figure 3-3** presents the XPS spectra of Co 2p core levels for  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  ( $x = 0.04, 0.07, 0.10$ ), where the fitting parameters are listed in **Table 3-2**. Since the Co 2p<sub>1/2</sub> and Co 2p<sub>3/2</sub> peaks could be divided into Co<sup>2+</sup> and Co<sup>3+</sup> peaks, the ratios of Co<sup>2+</sup>/(Co<sup>2+</sup>+Co<sup>3+</sup>) were estimated (**Table 3-3**). The Co<sup>2+</sup>/(Co<sup>2+</sup>+Co<sup>3+</sup>) ratio was increased with increasing  $x$ , indicating that more oxygen vacancies were generated. The charge transfer between Co<sup>2+</sup> and Co<sup>3+</sup> can accelerate the redox properties.



**Figure 3-3.** XPS spectra of Co 2p core levels for  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$ .

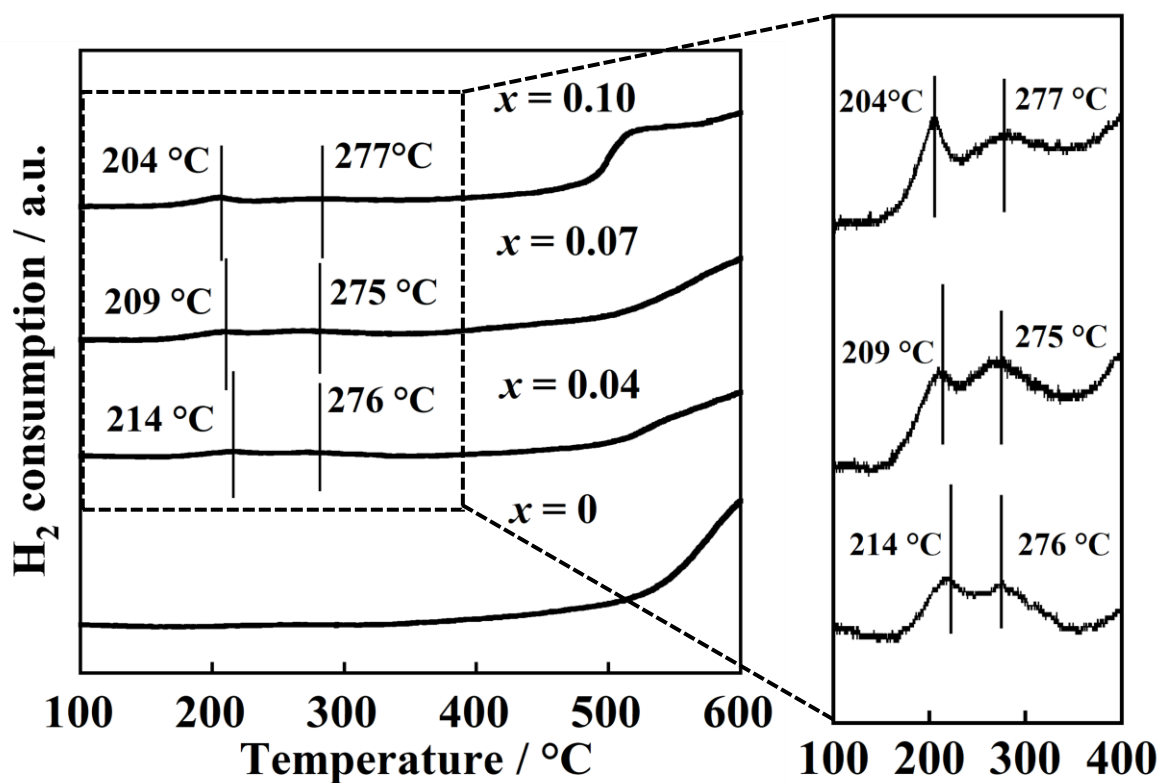
**Table 3-2.** Fitting parameters of the Co 2p XPS spectra for the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  catalysts

Catalyst	$\chi^2$	Peak	Position / eV	FWHM / eV	Area / %
$\text{Zr}_{1.09}\text{Sn}_{0.87}\text{Co}_{0.04}\text{O}_{4-\delta}$ ( $x = 0.04$ )	0.128	$\text{Co}^{3+} 2p_{3/2}$	773.3	7.3	13.2
		$\text{Co}^{2+} 2p_{3/2}$	780.8	4.4	14.1
		Satellite	786.1	7.5	44.9
		$\text{Co}^{3+} 2p_{1/2}$	790.9	7.8	9.3
		$\text{Co}^{2+} 2p_{1/2}$	796.5	4.9	5.3
		Satellite	802.9	7.8	13.1
$\text{Zr}_{1.13}\text{Sn}_{0.80}\text{Co}_{0.07}\text{O}_{4-\delta}$ ( $x = 0.07$ )	0.061	$\text{Co}^{3+} 2p_{3/2}$	773.3	7.3	15.2
		$\text{Co}^{2+} 2p_{3/2}$	780.7	4.5	19.8
		Satellite	786.0	7.6	32.4
		$\text{Co}^{3+} 2p_{1/2}$	790.9	7.8	6.9
		$\text{Co}^{2+} 2p_{1/2}$	796.5	4.8	10.1
		Satellite	802.8	7.7	15.6
$\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$ ( $x = 0.10$ )	0.019	$\text{Co}^{3+} 2p_{3/2}$	773.3	7.2	13.7
		$\text{Co}^{2+} 2p_{3/2}$	780.7	4.4	21.6
		Satellite	786.0	7.0	31.7
		$\text{Co}^{3+} 2p_{1/2}$	790.9	7.8	6.7
		$\text{Co}^{2+} 2p_{1/2}$	796.5	4.7	10.3
		Satellite	802.8	7.9	16.0

**Table 3-3.** Ratio of  $\text{Co}^{2+}/(\text{Co}^{2+}+\text{Co}^{3+})$  for  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$ 

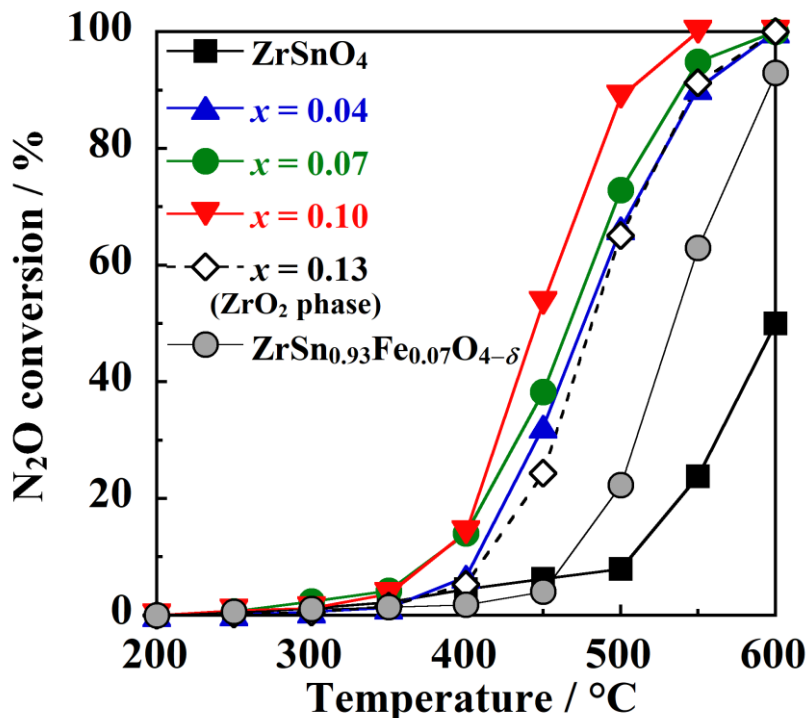
Catalyst	$\text{Co}^{2+}/(\text{Co}^{2+}+\text{Co}^{3+})$
$x = 0.04$	51.6 %
$x = 0.07$	56.6 %
$x = 0.10$	61.2 %

**Figure 3-4** shows the TPR profiles of  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  ( $x \leq 0.10$ ). For  $\text{ZrSnO}_4$ , there were no obvious reduction peaks below 300 °C, whereas the shoulder peak was observed at the temperatures higher than 300 °C. The  $\text{Co}^{2+/3+}$  doped samples had two reduction peaks below 300 °C, identified as the reduction of surface  $\text{Co}^{3+}$  and bulk  $\text{Co}^{3+}$  [42]. While no remarkable change was observed for the bulk reduction peak (ca. 275 °C) because of its low intensity and broad shape, the surface reduction peak (ca. 210 °C) decreased obviously with increasing  $x$ . These behaviors were attributed to the improved redox properties. The highest oxygen supply ability was obtained for  $\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$ .



**Figure 3-4.** TPR profiles of the  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  catalysts.

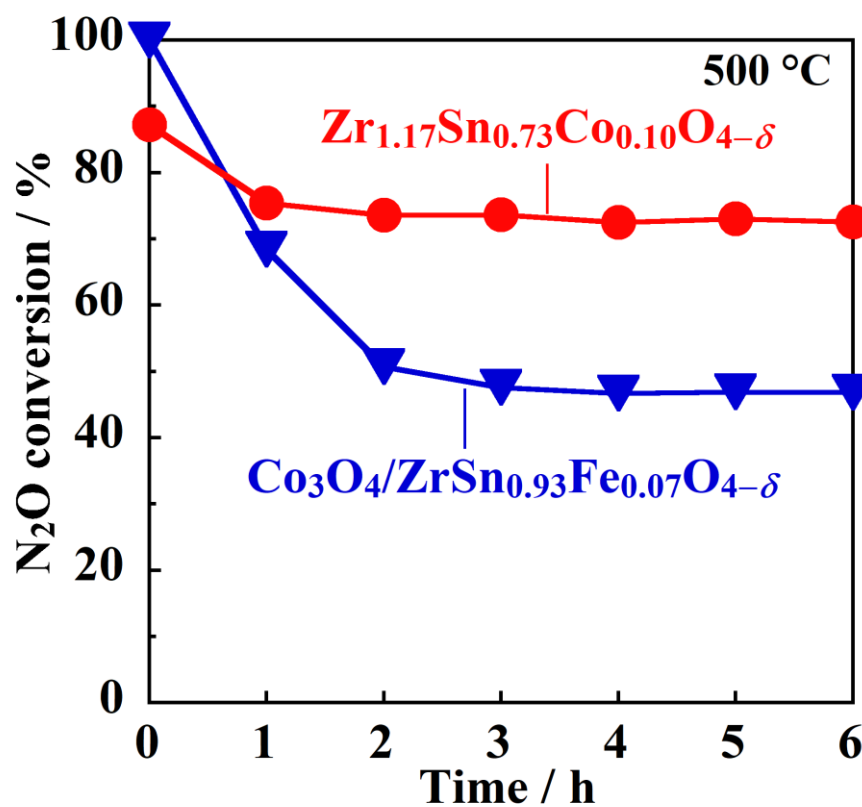
**Figure 3-5** shows the temperature dependence of the  $\text{N}_2\text{O}$  conversion for  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  with the data of  $\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$ . With increasing the Co content ( $x$ ), the catalytic activity increased, and the highest activity was obtained for the  $x = 0.10$  catalyst, in which  $\text{N}_2\text{O}$  was completely decomposed at 550 °C. This enhanced activity was attributed to the high oxygen supply ability that facilitated the removal of adsorbed oxygen, which is known to be the rate-determining step, as described in the General Introduction section. In addition, oxygen vacancies formed by replacing the  $\text{Sn}^{4+}$  site for lower-valent  $\text{Co}^{2+/3+}$  could act as catalytically active sites. In contrast, the  $x = 0.13$  catalyst with the cubic  $\text{ZrO}_2$  phase showed low activity compared to that of the  $x = 0.10$  catalyst with the  $\alpha$ - $\text{PbO}_2$ -type phase. Moreover, the  $\text{Co}^{2+/3+}$  doped catalyst of  $\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$  exhibited higher catalytic activity than that of the  $\text{Fe}^{2+/3+}$  doped catalyst ( $\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$ ).



**Figure 3-5.** Temperature dependence of  $\text{N}_2\text{O}$  conversion for  $\text{ZrSn}_{1-x}\text{Co}_x\text{O}_{4-\delta}$  and  $\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$ .



The durability test against the coexisting gases of oxygen gas and water vapor was investigated. **Figure 3-6** shows the N<sub>2</sub>O conversion in the presence of 10 vol% oxygen gas and 3 vol% water vapor at 500 °C as a function of time for the  $x = 0.10$  catalyst (Zr<sub>1.17</sub>Sn<sub>0.73</sub>Co<sub>0.10</sub>O<sub>4-δ</sub>), with the data of Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub>. Although Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> completely decomposed N<sub>2</sub>O before the durability test, it was drastically deactivated by the exposure of O<sub>2</sub> and H<sub>2</sub>O vapor. After 6 h, the N<sub>2</sub>O conversion for Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> decreased down to 47%; that is, the decrease in the N<sub>2</sub>O conversion on going from 0 to 6 h was 53% (100% → 47%). For Zr<sub>1.17</sub>Sn<sub>0.73</sub>Co<sub>0.10</sub>O<sub>4-δ</sub>, while the N<sub>2</sub>O conversion before the test (87%) was lower than that of Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> (100%), the high catalytic activity maintained even in the presence of O<sub>2</sub> and H<sub>2</sub>O vapor, where the slight decrease might be affected by the competition of coexisting gases with N<sub>2</sub>O for adsorption on the catalytic active sites. The decrease in the N<sub>2</sub>O conversion on going from 0 to 6 h was 14% (87% → 73%), which is appreciably lower than that for Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> (53%). Thus, the Co<sup>2+/3+</sup> doping was effective to obtain the catalyst with high durability, compared to the Co<sup>2+/3+</sup> loading as Co<sub>3</sub>O<sub>4</sub> with low durability. It is reported that Co<sub>3</sub>O<sub>4</sub> was deactivated by the adsorption of O<sub>2</sub> and/or H<sub>2</sub>O molecules on the surface Co<sup>2+/3+</sup> ions [43]. In the case of Zr<sub>1.17</sub>Sn<sub>0.73</sub>Co<sub>0.10</sub>O<sub>4-δ</sub>, since the Co<sup>2+/3+</sup> ions were introduced into the crystal lattice, the deactivation related to the surface Co species could be suppressed. Furthermore, the N<sub>2</sub>O conversion after 6 h for Zr<sub>1.17</sub>Sn<sub>0.73</sub>Co<sub>0.10</sub>O<sub>4-δ</sub> (73%) was considerably higher than that for Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> (47%).



**Figure 3-6.** N<sub>2</sub>O conversion in the presence of 10 vol% oxygen gas and 3 vol% water vapor as a function of time for  $\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$  and  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$ .

### 3.4. Conclusion

In this chapter, in order to develop novel catalysts with high durability against coexisting gases of O<sub>2</sub> and H<sub>2</sub>O vapor, the doping of Co<sup>2+/3+</sup> into the ZrSnO<sub>4</sub> lattice was conducted, instead of the loading of Co<sup>2+/3+</sup> as Co<sub>3</sub>O<sub>4</sub> with low durability; i.e., ZrSn<sub>1-x</sub>Co<sub>x</sub>O<sub>4-δ</sub> catalysts were prepared. The introduction of Co<sup>2+/3+</sup> into the ZrSnO<sub>4</sub> lattice resulted in the enhanced catalytic activity for direct decomposition of N<sub>2</sub>O. The Zr<sub>1.17</sub>Sn<sub>0.73</sub>Co<sub>0.10</sub>O<sub>4-δ</sub> catalyst exhibited the highest activity and achieved complete decomposition of N<sub>2</sub>O at 550 °C. Furthermore, the catalyst of Zr<sub>1.17</sub>Sn<sub>0.73</sub>Co<sub>0.10</sub>O<sub>4-δ</sub> showed high durability against coexisting gases of 10 vol% oxygen and 3 vol% water vapor. The decrease in the N<sub>2</sub>O conversion at 500 °C on going from 0 to 6 h was 14%, which was appreciably lower than that for Co<sub>3</sub>O<sub>4</sub>/ZrSn<sub>0.93</sub>Fe<sub>0.07</sub>O<sub>4-δ</sub> (53%).

## ***Summary***

In this doctoral thesis, novel catalysts based on zirconium-tin oxide were developed for effective direct decomposition of nitrous oxide. The results obtained through this study are summarized as follows:

### **Chapter 1**

In order to develop novel catalysts with high activity,  $\text{ZrSnO}_4$  was focused as a promoter to facilitate  $\text{N}_2\text{O}$  decomposition over the Rh activator. Among the  $x$  wt% Rh/ $\text{ZrSnO}_4$  catalysts, the 1.0 wt% Rh/ $\text{ZrSnO}_4$  catalyst exhibited the highest activity and was able to completely decompose  $\text{N}_2\text{O}$  at 450 °C. Furthermore,  $\text{ZrSnO}_4$  showed the higher oxygen supply ability than those of  $\text{SnO}_2$ ,  $\text{ZrO}_2$ , and  $\text{Al}_2\text{O}_3$ , likely because of the redox properties through the valence change of  $\text{Sn}^{4+/2+}$  and the oxygen vacancies for oxide ion migration formed by the coexistence of  $\text{Sn}^{2+}$ . The high oxygen supply ability of  $\text{ZrSnO}_4$  contributed to the high catalytic activity for 1.0 wt% Rh/ $\text{ZrSnO}_4$  compared to those of 1.0 wt% Rh/ $\text{SnO}_2$ , 1.0 wt% Rh/ $\text{ZrO}_2$ , and 1.0 wt% Rh/ $\text{Al}_2\text{O}_3$ .

### **Chapter 2**

For the development of noble-metal-free catalysts, spinel-type  $\text{Co}_3\text{O}_4$  was selected as an activator instead of Rh. Furthermore,  $\text{Fe}^{2+/3+}$  ions were introduced into the  $\text{ZrSnO}_4$  lattice to facilitate  $\text{N}_2\text{O}$  decomposition over the  $\text{Co}_3\text{O}_4$  activator. The  $\text{Fe}^{2+/3+}$  introduction improved the oxygen supply

ability, likely due to the redox cycle ( $\text{Fe}^{2+/3+}$ ) and the smooth oxide ion conduction via the oxygen vacancies generated by the replacement of the  $\text{Sn}^{4+}$  site for lower-valent  $\text{Fe}^{2+/3+}$ . Among the prepared catalysts, 19 wt%  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  exhibited the highest activity and achieved complete decomposition of  $\text{N}_2\text{O}$  at 500 °C without noble metal.

### Chapter 3

In order to develop novel catalysts with high durability against coexisting gases of  $\text{O}_2$  and  $\text{H}_2\text{O}$  vapor, the  $\text{Co}^{2+/3+}$  doping was carried out into the  $\text{ZrSnO}_4$  lattice, instead of the  $\text{Co}^{2+/3+}$  loading as  $\text{Co}_3\text{O}_4$  with low durability. The  $\text{Co}^{2+/3+}$  doping enhanced the oxygen supply ability, due to the improved redox properties caused by  $\text{Co}^{2+/3+}$  and the increased amount of oxygen vacancies owing to lower-valent  $\text{Co}^{2+/3+}$  compared to  $\text{Sn}^{4+}$ , which contributed to the high catalytic activity. Among the prepared catalysts, the  $\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$  catalyst exhibited the highest activity and achieved complete decomposition of  $\text{N}_2\text{O}$  at 550 °C. Moreover, the  $\text{Zr}_{1.17}\text{Sn}_{0.73}\text{Co}_{0.10}\text{O}_{4-\delta}$  catalyst showed high durability against coexisting gases of 10 vol% oxygen gas and 3 vol% water vapor; that is, the decrease in the  $\text{N}_2\text{O}$  conversion at 500 °C on going from 0 to 6 h was 14%, lower than that for  $\text{Co}_3\text{O}_4/\text{ZrSn}_{0.93}\text{Fe}_{0.07}\text{O}_{4-\delta}$  (53%).

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